

(19)



Europäisches Patentamt

European Patent Office

Office européen des brevets



(11)

EP 0 614 399 B1

(12)

EUROPEAN PATENT SPECIFICATION

(45) Date of publication and mention
of the grant of the patent:
08.03.2000 Bulletin 2000/10

(51) Int. Cl.⁷: **B01D 53/94**, B01J 23/10,
F24C 15/20

(21) Application number: 93900592.2

(86) International application number:
PCT/US92/10060

(22) Date of filing: 20.11.1992

(87) International publication number:
WO 93/10886 (10.06.1993 Gazette 1993/14)

(54) CERIA-ALUMINA OXIDATION CATALYST AND METHOD OF USE

CEROXID-ALUMINIUMOXID ENTHALTENDER OXIDATIONSKATALYSATOR UND VERFAHREN
ZUR ANWENDUNG.

CATALYSEUR PAR OXYDATION RECOURANT A L'OXYDE CERIQUE ET A L'OXYDE
D'ALUMINIUM ET PROCEDE D'UTILISATION

(84) Designated Contracting States:
AT BE CH DE DK ES FR GB GR IE IT LI LU MC NL
SE

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(30) Priority: 26.11.1991 US 798437
19.11.1992 US 973461

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(43) Date of publication of application:
14.09.1994 Bulletin 1994/37

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Description

BACKGROUND OF THE INVENTION5 Field of the Invention

[0001] This invention relates to a catalyst composition and method for the oxidation of oxidizable components of a gas-borne stream, e.g., for the treatment of diesel engine exhaust, and more specifically to the treatment of such diesel exhaust to reduce the particulates content thereof.

10 Background and Related Art

[0002] As is well-known, gas-borne streams or engine exhausts often contain oxidizable pollutants such as unburned fuel and vaporized or condensed oils. For example, diesel engine exhaust contains not only gaseous pollutants such as carbon monoxide ("CO") and unburned hydrocarbons ("HC"), but also soot particles which, as described in more detail below, comprise both a dry carbonaceous fraction and a hydrocarbon liquid which is sometimes referred to as a volatile organic fraction ("VOF"), which terminology will be used herein, or a soluble organic fraction. Accordingly, although sometimes loosely referred to as an "exhaust gas", the exhaust of a diesel engine is actually a heterogeneous material, comprising gaseous, liquid and solid components. The VOF may exist in diesel exhaust either as a vapor or as an aerosol (fine droplets of liquid condensate) depending on the temperature of the diesel exhaust.

[0003] Oxidation catalysts comprising a platinum group metal dispersed on a refractory metal oxide support are known for use in treating the exhaust of diesel engines in order to convert both HC and CO gaseous pollutants and particulates, i.e., soot particles, by catalyzing the oxidation of these pollutants to carbon dioxide and water. One problem faced in the treatment of diesel engine exhaust is presented by the presence of sulfur in diesel fuel. Upon combustion, sulfur forms sulfur dioxide and the oxidation catalyst catalyzes the SO_2 to SO_3 ("sulfates") with subsequent formation of sulfuric acid. The sulfates also react with activated alumina supports to form aluminum sulfates, which render activated alumina-containing catalysts inactive. In this regard, see U.S. Patent 4,171,289 at column 1, line 39 et seq. Previous attempts to deal with the sulfation problem include the incorporation of large amounts of sulfate-resistant materials such as vanadium oxide into the support coating, or the use of alternative support materials such as α -alumina, silica and titania, which are sulfation-resistant materials. Further, as is known, the oxidation of SO_2 to SO_3 also adds to the particulates in the exhaust by forming condensible sulfur compounds, such as sulfuric acid, which condense upon, and thereby add to, the mass of particulates.

[0004] Generally, the prior art has attempted to deal with these problems by dispersing a suitable oxidation catalyst metal, such as one or more platinum group metals, upon a refractory metal oxide support which is resistant to sulfation.

35 [0005] Examples of catalysts designed for the treatment of diesel exhaust fumes and soot include U.S. Patent 4,849,399 to Joy et al dated July 18, 1989. This Patent discloses catalytic composites which incorporate sulfur-resistant refractory inorganic oxides selected from the group consisting of titania, zirconia, and alumina treated with titania and/or zirconia (see column 6, lines 62-68).

[0006] U.S. Patent 4,759,918 to Homeier et al dated July 26, 1988 discloses catalysts for the treatment of diesel exhaust fumes and soot which incorporate sulfur-resistant refractory inorganic oxides selected from a group which includes silica, alumina, and silica-alumina (see column 3, lines 16-27).

40 [0007] EP-A-251,708 discloses certain catalysts of the type commonly referred to as "three-way conversion" ("TWC") catalysts, which catalyze simultaneously the oxidation of hydrocarbons, the oxidation of carbon monoxide and the reduction of nitrogen oxides. Such TWC catalysts find utility in the treatment of the exhaust gases from internal combustion engines fuelled by gasoline. The catalyst compositions with which this reference is concerned comprise a platinum-group metal catalytic component and an activated alumina that is stabilised against thermal degradation. Such a catalyst may be made by a method comprising the steps of applying a coating of activated alumina to a carrier substrate; calcining the resultant alumina-coated substrate to provide a calcined coating of activated alumina thereon; dispersing one or more platinum group metal components on the activated alumina; thereafter dispersing a stabilizer-precursor on the calcined coating on the carrier; and calcining the coating having the stabilizer-precursor dispersed thereon. The catalytic component may be selected from platinum, palladium and rhodium, and the stabilizer-precursor may be selected from precursors of stabilizers selected from alkaline earth metal (e.g. barium) oxide, silicon dioxide, rare earth metal (e.g. lanthanum) oxide and a Group IVB metal (e.g. zirconium) oxide. There is no disclosure in EP-A-251708 of a catalytic material that consists essentially of a combination of bulk ceria and a bulk alumina. Furthermore, 55 it is clear from the disclosure of this reference that the catalytic function is provided by the platinum-group metal component: there is no disclosure or suggestion that the combination of ceria and alumina as such may be relied upon for catalytic activity. Moreover, there is no disclosure in this reference of the use of a catalytic material consisting essentially of ceria and alumina to oxidize the volatile organic component of diesel exhaust gases. In addition, there is no disclo-

sure in this reference that a minor amount of a precious metal catalyst, in particular platinum or palladium, can be added to the ceria and alumina and will reduce the unwanted oxidation of SO_2 to SO_3 .

SUMMARY OF THE INVENTION

[0008] Generally, in accordance with the present invention, there is provided an oxidation catalyst composition and a method for oxidizing oxidizable components of a gas-borne stream, e.g., for treating diesel engine exhaust in which at least a volatile organic fraction component (described below) of the diesel exhaust particulates is converted to innocuous materials, and in which gaseous HC and CO pollutants may also be similarly converted. The objectives of the invention are attained by an oxidation catalyst comprising a base metal oxide catalytic material consisting essentially of a mixture of high surface area ceria and high surface area alumina, which optionally may have dispersed thereon a low loading of platinum catalytic metal. The method of the invention is attained by flowing a gas-borne stream, e.g., a diesel engine exhaust, into contact under reaction conditions with a catalyst composition as described above. In the case of treating diesel exhaust, the exhaust may be contacted under reaction conditions with a catalyst composition which contains palladium instead of a low loading of platinum but is otherwise as described above.

[0009] Specifically, in accordance with the present invention there is provided an oxidation catalyst composition which comprises a refractory carrier on which is disposed a coating of a ceria-alumina catalytic material consisting essentially of a combination of bulk ceria and bulk alumina each having a BET surface area of at least about $10 \text{ m}^2/\text{g}$, preferably the alumina having a surface area of from about $25 \text{ m}^2/\text{g}$ to $200 \text{ m}^2/\text{g}$ and the ceria having a surface area of from about $25 \text{ m}^2/\text{g}$ to $200 \text{ m}^2/\text{g}$.

[0010] In one embodiment of the invention, the ceria and alumina each comprises from about 5 to 95 percent, preferably from about 10 to 90 percent, more preferably from about 40 to 60 percent, by weight of the combination.

[0011] One aspect of the invention provides that the catalyst composition optionally further comprises a catalytically effective amount of platinum dispersed on the catalytic material in an amount not to exceed about 15 g/ft^3 (530 g/m^3) of the catalyst composition. For example, the platinum may be present in the amount of from about 0.1 to 15 g/ft^3 (3.5 to 530 g/m^3) of the composition, preferably from about 0.1 to 5 g/ft^3 (3.5 to 177 g/m^3) of the composition. When the catalyst composition includes platinum, another aspect of the invention provides that at least a catalytically effective amount of the platinum is dispersed on the ceria. At least a catalytically effective amount of the platinum may also be dispersed on the alumina. Such dispersal of the platinum may be utilized whether the alumina and ceria are mixed in a single layer or are present in discrete layers of, respectively, ceria and alumina and, in the latter case, irrespective of which of the two layers is the top layer.

[0012] Still another aspect of the invention provides that the ceria comprises an aluminum-stabilized ceria. The alumina may also be stabilized against thermal degradation. The ceria and alumina may be combined as a mixture and the mixture deposited as a single layer coating on the refractory carrier, or the ceria and alumina may be present in respective discrete superimposed layers of ceria and alumina. The ceria layer may be above or below the alumina layer. In accordance with the method of the present invention, there is provided a method of treating diesel engine exhaust containing a volatile organic fraction. The method includes contacting the exhaust with a catalyst composition comprised of components as described above or with a catalyst composition comprised of components as described above but which optionally includes palladium instead of the optional platinum. Thus, the method includes contacting the gas-borne stream to be treated with a catalyst composition comprising ceria and alumina as described above, and optionally including platinum or palladium. When the optional palladium is employed in the composition, it may be present in the amount from about 0.1 to 200 g/ft^3 (3.5 to 706 g/m^3), preferably in the amount of from about 20 to 120 g/ft^3 (706 to 424 g/m^3), of the catalyst composition. In accordance with the method of the present invention, contacting of the diesel exhaust with the catalyst composition is carried out at a temperature high enough to catalyze oxidation of at least some of the volatile organic fraction of the exhaust, for example, an inlet temperature of from about 100°C to 800°C .

DEFINITIONS

[0013] As used herein and in the claims, the following terms shall have the indicated meanings.

[0014] The term "gas-borne stream" means a gaseous stream which may contain non-gaseous components such as solid particulates and/or vapors, liquid mist or droplets, and/or solid particulates wetted by a liquid.

[0015] The term "BET surface area" has its usual meaning of referring to the Brunauer, Emmett, Teller method for determining surface area by N_2 adsorption. Unless otherwise specifically stated, all references herein to the surface area of a ceria, alumina or other component refer to the BET surface area.

[0016] The term "activated alumina" has its usual meaning of a high BET surface area alumina, comprising primarily one or more of γ -, θ - and δ -aluminas (gamma, theta and delta).

[0017] The term "catalytically effective amount" means that the amount of material present is sufficient to affect the rate of reaction of the oxidation of pollutants in the exhaust being treated.

[0018] The term "inlet temperature" shall mean the temperature of the exhaust, test gas or other stream being treated immediately prior to initial contact of the exhaust, test gas or other stream with the catalyst composition.

[0019] The term "ceria-alumina catalytic material" means a combination of ceria particles and alumina particles each having a BET surface area of at least about 10 m²/g, i.e., a combination of high surface area bulk ceria and high surface area bulk alumina, sometimes referred to as "activated alumina".

[0020] The term "combination" when used with reference to a combination of ceria and alumina includes combinations attained by mixtures or blends of ceria and alumina as well as superimposed discrete layers of ceria and alumina.

[0021] The term "aluminum-stabilized ceria" means ceria which has been stabilized against thermal degradation by incorporation therein of an aluminum compound. A suitable technique is shown in U.S. Patent 4,714,694 of C.Z. Wan et al, in which ceria particles are impregnated with a liquid dispersion of an aluminum compound, e.g., an aqueous solution of a soluble aluminum compound such as aluminum nitrate, aluminum chloride, aluminum oxychloride, aluminum acetate, etc. After drying and calcining the impregnated ceria in air at a temperature of, e.g., from about 300°C to 600°C for a period of 1/2 to 2 hours, the aluminum compound impregnated into the ceria particles is converted into an effective thermal stabilizer for the ceria. The term "aluminum-stabilized" is used for economy of expression although the aluminum is probably present in the ceria as a compound, presumably alumina, and not as elemental aluminum.

[0022] Reference herein or in the claims to ceria or alumina being in "bulk" form means that the ceria or alumina is present as discrete particles (which may be, and usually are, of very small size, e.g., 10 to 20 microns (μm) in diameter or even smaller) as opposed to having been dispersed in solution form into another component. For example, the thermal stabilization of ceria particles (bulk ceria) with alumina as described above with respect to U.S. Patent 4,714,694 results in the alumina being dispersed into the ceria particles and does not provide the dispersed alumina in "bulk" form, i.e., as discrete particles of alumina.

[0023] The abbreviation "TGA" stands for thermogravimetric analysis which is measure of the weight change (e.g., loss) of a sample as a function of temperature and/or time. The abbreviation "DTA" stands for differential thermal analysis which is measure of the amount of heat emitted (exotherm) or absorbed (endotherm) by a sample as a function of temperature and/or time.

BRIEF DESCRIPTION OF THE DRAWINGS

[0024]

Figure 1 is a plot of oxidation of SO₂ to SO₃ in a gas stream being treated with an oxidation catalyst, the degree of oxidation being plotted on the ordinate versus the platinum loading of the catalyst on the abscissa;

Figure 2 is a plot similar to that of Figure 1 but showing the degree of HC oxidation on the ordinate versus platinum loading on the abscissa;

Figure 3 is a perspective plot of oxidation of SO₂ to SO₃ in a gas stream being treated with an oxidation catalyst, with the degree of oxidation indicated by the height of the vertical bars for three different samples, each containing 0.5 g/ft³ (18 g/m³) of platinum and having different weight percentages of ceria in the ceria-alumina catalytic material;

Figure 4 is a plot of a factor (DTA peak area) correlating combustion of engine lubricating oil (simulating the unburned lubricating oil in "VOF", described below) plotted on the ordinate versus the platinum content of a ceria-alumina washcoat used to catalyze the combustion of the lubricating oil plotted on the abscissa; and

Figures 5 through 8 are plots showing various aspects of diesel engine exhaust treatment performance of three aged catalyst samples made in accordance with certain embodiments of the present invention as a function of the operating temperature of the catalysts, as follows: Figure 5 shows the percentage conversion of the volatile organic fraction ("VOF"); Figure 6 shows the percentage conversion of total particulate matter ("TPM") in the exhaust; Figure 7 shows the gas phase conversion of hydrocarbons ("HC") and Figure 8 shows the gas phase conversion of carbon monoxide ("CO").

DETAILED DESCRIPTION OF THE INVENTION AND PREFERRED EMBODIMENTS THEREOF

[0025] The present invention provides an oxidation catalyst composition which is effective for oxidizing oxidizable components of a gas-borne stream, for example, for treating diesel engine exhaust. In the latter case, the composition is particularly effective with regard to reducing the total particulates in the exhaust. The carbonaceous particulates ("soot") component of diesel engine exhaust is, as is well-known, comprised of two major components. One component is relatively dry carbonaceous particles and the other, usually referred to as a volatile organic fraction ("VOF"), is a mixture of high molecular weight hydrocarbons comprised of unburned and partially burned diesel fuel and lubricating oil. The volatile organic fraction is present in the diesel exhaust as either a vapor phase or a liquid phase, or both, depending on the temperature of the exhaust. Generally, it is not feasible to attempt to remove or treat the dry, solid carbonaceous

aceous particulates component of the total particulates by catalytic treatment, and it is the VOF component which can be most effectively removed by conversion via utilization of an oxidation catalyst. Therefore, in order to reduce the total particulates discharged so as to meet present and impending Government regulations concerning maximum allowable total particulates, the volatile organic fraction, or at least a portion thereof, is oxidized to innocuous CO_2 and H_2O by being contacted with an oxidation catalyst under suitable reaction conditions. The required U.S. Government limits for 1991 on HC, CO, nitrogen oxides (" NO_x ") and total particulate emissions ("TPM") in diesel engine exhaust have been largely met by suitable engine design modifications. For 1994 the HC, CO and NO_x limits remain unchanged from 1991 standards but the upper limit on TPM will be reduced from the 1991 level of 0.25 grams per horsepower-hour ("g/HP-hr") to 0.10 g/HP-hr (0.93 to 0.037 g/MJ). Although the oxidation catalysts of the present invention, when employed as a diesel exhaust treatment catalyst, are primarily concerned with effectuating a reduction in total particulates, they are also capable, with the optional addition of platinum in limited amounts of providing the added advantage of also oxidizing a portion of the HC and CO contained in the gaseous component of the diesel engine exhaust without promoting excessive oxidation of SO_2 to SO_3 . The oxidation catalysts of the present invention avoid or reduce the unwanted side effect of promoting the oxidation of SO_2 to SO_3 which, as noted above, contributes to the particulates problem because the condensation of sulfuric acid and other sulfate condensibles which accumulate on, and add to, the mass of the particulates in the exhaust.

[0026] However, the oxidation catalysts of the present invention have utility for uses other than the treatment of diesel engine exhaust. Generally, the catalysts of the present invention are useful for oxidation of gas-borne oxidizable components in engine exhausts generally, such as any application in which lubricating oils are discharged, e.g., the exhaust of compressed natural gas engines, ethanol-fueled engines, compressors, gas turbines, etc. Many alternate-fueled engines such as compressed natural gas engines are built on diesel engine carcasses and therefore inherently discharge significant quantities of lubricating oil.

[0027] In accordance with the teachings of the present invention it has been found, surprisingly, that the beneficial effect of oxidizing pollutants generally, and in particular of reducing diesel exhaust particulates emissions by oxidation of the volatile organic fraction thereof, can be attained by a mixture of high surface area, i.e., activated, alumina and a high surface area ceria, each having a BET surface area of $10 \text{ m}^2/\text{g}$ or higher. For purposes of illustration, the benefits of the present invention will be described in detail below with respect to the treatment of diesel engine exhaust. The basic and novel characteristics of the present invention are believed to reside in the use of the defined combination of ceria and alumina as an oxidation catalyst without the addition of metal catalytic components thereto, except as specifically otherwise defined in certain dependent claims. Preferably, the bulk ceria and the bulk alumina will each have a surface area of at least about $10 \text{ m}^2/\text{g}$, preferably at least about $20 \text{ m}^2/\text{g}$. For example, the bulk alumina may have a surface area of from about 120 to $180 \text{ m}^2/\text{g}$ and the bulk ceria may have a surface area of from about 70 to $150 \text{ m}^2/\text{g}$. The fact that a catalyst composition which can serve as a diesel oxidation catalyst and which contains activated alumina as a major component thereof has proven to be successful is in itself surprising, in view of the consensus of the prior art that alumina, if used at all in diesel oxidation catalysts, must be a low surface area alumina (α -alumina) and/or be used in conjunction with sulfate-resistant refractory metal oxides such as zirconia, titania or silica. It has nonetheless been found that, in accordance with the present invention, surprisingly, a combination of high surface area alumina and a high surface area ceria provides a catalytic material which effectively catalyzes the oxidation of the volatile organic fraction so as to provide a significant reduction in total particulates in diesel engine exhaust and exhibits good durability, that is, long life, both in laboratory and diesel engine tests. It should be noted that the prior art generally considers refractory base metal oxides used in diesel oxidation catalysts to be merely supports for the dispersal thereon of catalytically active metals such as platinum group metals. In contrast, the present invention teaches that a ceria-alumina catalytic material comprising essentially only ceria and alumina of sufficiently high surface area ($10 \text{ m}^2/\text{g}$ or higher); dispersed on a suitable carrier, provides a durable and effective diesel oxidation catalyst.

[0028] It has further been found that beneficial effects are attained by the optional incorporation of platinum in the catalyst composition, provided that the platinum is present at loadings much lower than those conventionally used in oxidation catalysts. It has been discovered that, most surprisingly, a limited quantity of platinum in the catalyst composition actually reduces the undesirable oxidation of SO_2 to SO_3 relative to that encountered by using the ceria-alumina catalytic material alone, while nonetheless promoting some oxidation of CO and HC gaseous components of the diesel exhaust. The suppression of the oxidation of SO_2 to SO_3 by the addition of low loadings of platinum is a very surprising finding, given the powerful catalytic activity of platinum in promoting oxidation reactions generally. Without wishing to be bound by any particular theory, it may be that the presence of a low loading of platinum on the ceria occupies some catalytic sites on the ceria, thereby moderating the tendency of ceria to promote the oxidation of SO_2 to SO_3 . If the catalytic metal platinum is added to the catalytic composition, it serves to catalyze the oxidation of gas phase HC and CO pollutants as an added benefit. However, such catalytic metal is not needed to supplement the action of the ceria-alumina catalytic material in reducing total particulate emissions. The platinum catalytic metal does not appear to play a role in controlling particulates, as indicated by data discussed elsewhere herein, which show that the quantity of platinum utilized does not significantly affect the rate of particulates conversion.

[0029] The catalysts of the present invention may take the form of a carrier or substrate, such as a monolithic "honeycomb" structure (a body having a plurality of gas flow passages extending therethrough), on which is applied a coating of the catalytic material comprising a mixture of high surface area ceria and alumina and, optionally, a low loading platinum. As discussed below, discrete coatings of the ceria and alumina may be employed.

The Carrier (Substrate)

[0030] The carrier used in this invention should be relatively inert with respect to the catalytic composition dispersed thereon. The preferred carriers are comprised of ceramic-like materials such as cordierite, α -alumina, silicon nitride, zirconia, mullite, spodumene, alumina-silica-magnesia or zirconium silicate, or of refractory metals such as stainless steel. The carriers are preferably of the type sometimes referred to as honeycomb or monolithic carriers, comprising a unitary cylindrical body having a plurality of fine, substantially parallel gas flow passages extending therethrough and connecting both end-faces of the carrier to provide a "flow-through" type of carrier. Such monolithic carriers may contain up to about 700 or more flow channels ("cells") per square inch of cross section, although far fewer may be used. For example, the carrier may have from about 7 to 600, more usually from about 200 to 400, cells per square inch ("cps") (108.5 to 9 300, more usually 3 100 to 6 200, cells per dm²).

[0031] While this discussion and the following examples relate to flow-through type carrier substrates, wall-flow carriers (filters) may also be used. Wall-flow carriers are generally similar in structure to flow-through carriers, with the distinction that each channel is blocked at one end of the carrier body, with alternate channels blocked at opposite end-faces. Wall-flow carrier substrates and the support coatings deposited thereon are necessarily porous, as the exhaust must pass through the walls of the carrier in order to exit the carrier structure.

The Catalytic Material

[0032] The ceria-alumina catalytic material may be prepared in the form of an aqueous slurry of ceria and alumina particles, the particles optionally being impregnated with the platinum catalytic metal component if one is to be utilized. The slurry is then applied to the carrier, dried and calcined to form a catalytic material coating ("washcoat") thereon. Typically, the ceria and alumina particles are mixed with water and an acidifier such as acetic acid, nitric acid or sulfuric acid, and ball milled to a desired particle size.

[0033] The optional platinum catalytic metal component is, when used, preferably incorporated into the ceria particles or into the ceria and alumina particles. In such case, the ceria-alumina acts not only as a catalyst but also as a support for the optional platinum catalytic metal component. Such incorporation may be carried out after the ceria-alumina catalytic material is coated as a wash-coat onto a suitable carrier, by impregnating the coated carrier with a solution of a suitable platinum compound, followed by drying and calcination. However, preferably, the ceria particles or both the ceria and alumina particles are impregnated with a suitable platinum compound before a coating of the ceria-alumina catalytic material is applied to the carrier. In either case, the optional platinum metal may be added to the ceria-alumina catalytic material as, e.g., a solution of a soluble platinum compound, the solution serving to impregnate the ceria and alumina particles (or the ceria-alumina coating on the carrier), which may then be dried and the platinum fixed thereon. Fixing may be carried out by calcination or by treatment with hydrogen sulfide or by other known means, to render the metal in water-insoluble form.

[0034] Generally, the slurry of ceria and activated alumina particles, whether or not impregnated with the platinum compound solution, will be deposited upon the carrier substrate and dried and calcined to adhere the catalytic material to the carrier and, when the platinum compound is present, to revert the platinum compound to the elemental metal or its oxide. Suitable platinum compounds for use in the foregoing process include potassium platinum chloride, ammonium platinum thiocyanate, amine-solubilized platinum hydroxide and chloroplatinic acid, as is well-known in the art. During calcination, or at least during the initial phase of use of the catalyst, such compounds, if present, are converted into the catalytically active elemental platinum metal or its oxide.

[0035] When the catalytic material is applied as a thin coating to a suitable carrier, such as described above, the proportions of ingredients are conventionally expressed as weight of material per unit volume of catalyst, as this measure accommodates the presence of different sizes of catalyst composition voids provided by different carrier wall thicknesses, gas flow passages, etc. Grams per cubic inch ("g/in³") units (1 g/in³ = 61.024 g/dm³) are used to express the quantity of relatively plentiful components such as the ceria-alumina catalytic material, and grams per cubic foot ("g/ft³") units (1 g/ft³ = 35.315 g/m³) are used to express the quantity of the sparsely used ingredients, such as the platinum metal. For typical diesel exhaust applications, the ceria-alumina catalytic material of the present invention generally may comprise from about 0.25 to about 4.0 g/in³ (15 to 244 g/dm³), preferably from about 0.25 to about 3.0 g/in³ (15 to 183 g/dm³) of the coated carrier substrate, optionally including from about 0 to 25 g/ft³ (0 to 883 g/m³), preferably from about 0 to 15 g/ft³ (0 to 530 g/m³) of platinum.

[0036] Without wishing to be bound by a particular theory, applicants offer the following hypothesis to explain the

superior performance, when used to treat diesel engine exhaust, of the ceria-alumina catalytic materials according to this invention. It is believed that diesel exhaust contains a significant proportion of gases or vapors which are close to their dew point, i.e., close to condensing to a liquid, and thereby adding to the VOF portion of the particulates at the conditions obtaining in the exhaust pipe. These "potential particulates" condense in the ceria-alumina catalytic materials, their condensation being enhanced by a capillary condensation effect, a known phenomenon in which a capillary-like action facilitates condensation of oil vapors to liquid phase. The small pore size of the high surface area ceria-alumina catalytic material is believed to provide such capillary condensation action for the VOF. Generally, the higher the surface area of the ceria and alumina, the smaller is their pore size. As the exhaust temperature increases during increased work loads imposed on the diesel engine, the condensed hydrocarbon liquids (condensed VOF) are desorbed from the ceria-alumina catalytic material and volatilize, at which time the catalytic effect of the ceria-alumina catalytic material, which provides numerous acidic sites, is believed to enhance cracking and gas phase oxidation, i.e., combustion, of the desorbed, re-volatilized hydrocarbon (VOF) vapors. Even if a proportion of the vapors re-volatilized from the condensate is not combusted, the cracking of heavy VOF components to lighter hydrocarbons reduces the total amount of condensibles, so that the total particulates output from the diesel engine is concomitantly further reduced. In this latter regard, the ceria-alumina catalytic material is believed to act as a trap and a storage medium for condensed or condensible VOF during relatively cool phases of the exhaust, and releases the cracked VOF only upon re-volatilization thereof during relatively hot phases. The porous nature of the ceria-alumina catalytic material is also believed to promote rapid diffusion of the VOF throughout the washcoat structure, thereby facilitating relatively low temperature gasification and oxidation of the VOF upon increases in temperature of the catalyst during higher engine load (and therefore increased exhaust gas temperature) cycles. Data on aging show that the presence of sulfates does not significantly adversely affect the capacity of the ceria-alumina catalytic material to reduce particulate emissions.

[0037] Generally, other ingredients may be added to the catalyst composition of the present invention such as conventional thermal stabilizers for the alumina, e.g., rare earth metal oxides such as ceria. Thermal stabilization of high surface area ceria and alumina to militate against phase conversion to less catalytically effective low surface area forms is well-known in the art although thermal stabilization of alumina is not usually needed for diesel exhaust service. Such thermal stabilizers may be incorporated into the bulk ceria or into the bulk activated alumina, by impregnating the ceria (or alumina) particles with, e.g., a solution of a soluble compound of the stabilizer metal, for example, an aluminum nitrate solution in the case of stabilizing bulk ceria. Such impregnation is then followed by drying and calcining the impregnated ceria particles to convert the aluminum nitrate impregnated therein into alumina.

[0038] In addition, the catalyst compositions of the invention may contain other catalytic ingredients such as other base metal promoters or the like. However, in one embodiment, the catalyst composition of the present invention consists essentially only of the high surface area ceria and high surface area alumina, preferably present in a weight proportion of 1.5:1 to 1:1.5, with or without thermal stabilizers impregnated therein, and, optionally, limited amounts of platinum. With respect to the method aspect of the invention, the use of palladium in place of platinum is contemplated.

Examples and Data

[0039] A catalyst composition in accordance with one embodiment of the invention, in which an optional alumina undercoat is provided beneath a coating of the ceria-alumina catalytic material having a platinum metal dispersed thereon, was prepared as follows.

Example 1

[0040] A. An activated alumina undercoat slurry is prepared by combining 1000 grams of activated alumina having a nominal BET surface area of 150 m²/g with 50 cubic centimeters ("cc") of glacial acetic acid and 1 cc of an anti-foamant sold under the trademark NOPCO NXZ in 1000 cc of deionized water. The ingredients are ball milled until an average particle size of at least 90 percent by volume of the particles having a diameter of not greater than 12 microns is attained. Cylindrical carriers comprising cordierite cylinders 6 inches (15.24 cm) long by 6 inches (15.24 cm) in diameter and having 400 gas flow passages per square inch of end face area (400 cpsi or 6 200 per dm²) are dipped into the slurry, excess slurry is blown from the gas flow passages and the slurry-coated carriers are dried at 110°C and then calcined in air at 450°C for 1 hour to provide alumina-coated carriers.

[0041] B. The ceria-alumina catalytic material is prepared by utilizing 1050 grams of the same activated alumina as used in Part A and 900 grams of aluminum-stabilized ceria having a BET surface area of 164 m²/g. The aluminum-stabilized ceria is attained by impregnating the ceria particles with a solution of an aluminum compound such as aluminum nitrate followed by calcining, to provide an aluminum content in the ceria of 1.35 weight percent aluminum, based on the total weight of ceria with the weight of aluminum calculated as the metal. Presumably, the aluminum is present as alumina. One such method of preparing an aluminum-stabilized ceria is shown in U.S. Patent 4,714,694 issued December 22, 1981 to C.Z. Wan et al. As is well-known, high surface area refractory oxides such as ceria are subject to loss

of surface area and consequent reduction in catalytic efficiency upon prolonged exposure to high temperatures and other conditions of treating diesel exhausts.

[0042] Aluminum-stabilized ceria is more resistant to such thermal degradation than is unstabilized ceria. As is also well-known, alumina may also be thermally stabilized, usually by a similar impregnation of the alumina with precursors of rare earth metal oxides such as ceria. However, thermal stabilization of the alumina is usually not necessary for the temperatures encountered in treating diesel engine exhaust. The high surface area ceria and high surface area alumina particles are placed in separate ball mills. A quantity of an amine-solubilized platinum hydroxide solution containing 0.2894 grams of platinum, a quantity of monoethanolamine ("MEA"), 97.5 cc of glacial acetic acid, 2.0 cc of an anti-foamant sold under the trademark NOPCO NXZ and about 1950 cc of deionized water are employed. About one-half the water and sufficient MEA to adjust the pH to at least about 7 are placed in the ball mill containing the alumina which is milled to thoroughly blend the ingredients. Then, one-half of the platinum solution is added and ball milling is continued for about 5 minutes. Thereafter, about one-half the glacial acetic acid and anti-foamant are added and milling is continued until a particle size of at least about 90 percent by weight of the particles having a diameter of less than about 12 microns is attained. The same process is separately repeated with the aluminum-stabilized ceria, except that MEA is not employed, including ball milling for mixing and to attain the same particle size of the ceria particles. The alumina and ceria slurries are then blended together to form a slurry of alumina and ceria particles containing a platinum compound. The alumina-coated carrier obtained in Part A of this Example 1 is dipped into the blended slurry, excess slurry is blown from the gas flow passages of the carrier, and the coated carrier is then dried and calcined in air at 450°C to provide a finished catalyst containing a coating of a ceria-alumina catalytic material having about 0.5 g/ft³ (18 g/m³) of platinum dispersed thereon. The catalytic material coating, sometimes referred to as a washcoat, inclusive of the platinum content, comprises about 1.95 g/in³ (119 g/dm³) of the catalyst composition, the catalytic material overlying an alumina undercoat which comprises about 1.00 g/in³ (61 g/dm³) of the catalyst composition. Unless otherwise specified, catalyst samples in accordance with the present invention in subsequent Examples have the same type and loading of alumina undercoat and ceria-alumina catalytic material as a topcoat overlying the undercoat.

[0043] Reference in the following TABLES, or elsewhere in this application, to a percentage conversion of constituents (rendered as "%C" in the TABLES) of the exhaust or test gas, means the percentage of such constituent initially present in the exhaust or test gas being treated which is converted to another species, e.g., the conversion to H₂O and/or CO₂ of HC, CO and VOF, and the oxidation to SO₃ of SO₂. Thus, if an exhaust contains 10 volume percent CO and treatment of the exhaust results in an outlet gas containing 6 volume percent CO, a 40 percent conversion of the CO has been attained. Reference in the following Examples, or elsewhere in this application, to "space velocity" means the flow rate of exhaust or test gas flowed through a catalyst, expressed as volumes of exhaust or test gas per volume of catalyst per hour, calculated with the exhaust or test gas at standard conditions of temperature and pressure.

Example 2

[0044] A series of sample catalysts was prepared generally in accordance with the procedures of Example 1 to provide a series of five otherwise identical compositions containing a ceria-alumina catalytic material in accordance with the teachings of the present invention, having various amounts of platinum dispersed thereon, including 0, 0.5, 1.0, 2.0 and 5.0 g/ft³ (0, 18, 35, 71 and 177 g/m³) of platinum. These catalyst samples comprised cores measuring 1.5 inches (3.81 cm) in diameter and 3.0 inches (7.62 cm) in length, cut from cordierite carriers 6 inches long and 6 inches in diameter, used, as in Example 1, to make the catalysts of this Example 2. The resulting 400 cpsi (6200 cells/dm²) cordierite sample cores contained a loading of 1.95 g/in³ (119 g/dm³) of the ceria-alumina catalytic material overlying an alumina undercoat present in the amount of 1.00 g/in³ (61 g/dm³), in addition to the specified loading of platinum metal dispersed on the ceria-alumina catalytic material. The test catalysts were aged for 10 hours at 500°C by having a mixture of 10 percent steam in air flowed through them. A test gas was contacted with each of these aged catalysts in a series of tests at a space velocity of 50,000 and inlet temperature of, respectively, 275°C, 350°C, 425°C and 500°C. The test gas had a composition of 10 percent steam, 10 percent oxygen, 4.5 percent CO₂, 1000 ppm NO, 28.6 ppm heptane, 200 ppm CO, 50 ppm SO₂, balance nitrogen. All percents are volume percent and "ppm" means parts per million by volume. Measurements were taken to determine the amount of oxidation of SO₂ to SO₃. The results of these tests are tabulated in TABLE I below and plotted in Figure 1.

TABLE I

Inlet Gas Temp. (°C)	Platinum Loading (g/ft ³) ^b	%C ^a SO ₂	%C ^a HC	%C ^a CO
275	0	8.0	0.0	0.0
275	0.5	0.0	2.4	30.5
275	1.0	6.1	0.0	74.6
275	2.0	16.0	10.0	99.0
275	5.0	30.6	20.2	99.5
350	0	8.0	0.0	5.9
350	0.5	4.0	9.8	68.3
350	1.0	17.6	31.7	97.9
350	2.0	21.6	87.8	100
350	5.0	30.0	83.1	100
425	0	12.0	2.6	10.3
425	0.5	11.8	31.6	84.3
425	1.0	25.5	66.6	96.4
425	2.0	33.3	90.5	100
425	5.0	48.0	91.9	100
500	0	20.0	9.3	9.3
500	0.5	12.0	47.4	84.8
500	1.0	28.8	80.5	98.5
500	2.0	35.3	83.1	99.5
500	5.0	62.0	88.0	100

^a %C^a means the percentage conversion of the indicated constituent.

^b 1 g/ft³ = 35.315 g/m³

[0045] The data of TABLE I, and the plot thereof in Figure 1, clearly show that the ceria-alumina catalytic material containing no platinum in each case provided, at each temperature level tested, a somewhat higher degree of conversion of SO₂ to SO₃ than did the otherwise identical ceria-alumina catalyst containing 0.5 g/ft³ (18 g/m³) of platinum. As the platinum loading was increased to 1.0 g/ft³ (35 g/m³), at each temperature level, the degree of undesired conversion of SO₂ to SO₃ increased as compared to the versions containing no or only 0.5 g/ft³ (18 g/m³). Further increases in platinum loading to 2 and 5 g/ft³ (71 and 177 g/m³) further increased, as one would expect, the oxidation of SO₂. What is very surprising is the fact, clearly shown in Figure 1 and the data of TABLE I, that the ceria-alumina catalytic material containing 0.5 g/ft³ (177 g/m³) of platinum dispersed thereon demonstrated less conversion of SO₂ to SO₃ than did the ceria-alumina catalytic material containing no platinum metal thereon. As noted above, it is believed that the presence of a low loading of platinum on the ceria may occupy some catalytic sites which otherwise are highly effective in promoting the oxidation of SO₂ to SO₃.

[0046] Figure 2 shows the corresponding conversion of hydro-carbons in the test gas at the various temperature levels tested. The HC and CO conversion data of TABLE I, and the plot of the HC conversion data of TABLE I in Figure 2, show the expected result that as the content of platinum metal increases the degree of conversion of HC and CO likewise increases. As discussed elsewhere herein, because of successful modifications in diesel engine design, catalytic treatment of diesel exhaust may not be necessary in order to attain reductions in HC and CO to meet U.S. Government standards, because the modified engines have reduced the output of HC and CO to below that of the current and impending U.S. Government standards. Nonetheless, the inclusion of platinum, at least at a loading of not more than about 1 g/ft³ (35 g/m³), preferably at from about 0.1 to 0.8 g/ft³ (3.5 to 28 g/m³), more preferably at about 0.5 g/ft³ (18 g/m³), is seen to have a beneficial effect on reducing the amount of oxidation of SO₂ to SO₃. Thus limited, the addition

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of platinum is seen to reduce SO₂ oxidation and thereby ameliorate particulates emissions. The addition of platinum also provides a beneficial added effect of further reducing HC and CO emissions.

[0047] It will be appreciated that in some cases it may be desired or necessary to significantly reduce HC and/or CO emissions and, in order to do so, the addition of moderate amounts of platinum, not more than 15 g/ft³ (530 g/m³), preferably not more than 5 g/ft³ (177 g/m³), and most preferably not more than 2 g/ft³ (71 g/m³), may be desirable despite the concomitant increase in SO₂ oxidation at additions of significantly more than 0.5 g/ft³ (18 g/m³).

Example 3

[0048] A series of test catalysts was prepared generally in accordance with the procedure outlined in Example 1 to provide three samples, each comprising an alumina undercoat at a loading of 1.0 g/in³ (61 g/dm³) upon which was coated a topcoat layer comprised of a ceria-alumina catalytic material containing ceria and alumina in proportions of 46.2 weight percent aluminum-stabilized ceria and 53.8 weight percent alumina, and having dispersed thereon 0.5 g/ft³ (18 g/m³) of platinum. The ceria-alumina topcoat layer was present in the amount of 1.95 g/in³ (119 g/dm³). The ceria had a surface area of about 164 m²/g and the alumina had a surface area of about 150 m²/g. One sample, designated S-3Ce, has the platinum dispersed only on the ceria component of the catalytic material, a second sample designated S-3 has equal amounts of the platinum dispersed on the ceria and the alumina components of the ceria-alumina catalytic material, and the third sample, designated S-3Al, has the platinum disposed entirely on the alumina component of the ceria-alumina catalytic material. The three catalyst samples were then tested for HC, CO and SO₂ conversion at 350°C and a space velocity of 90,000. The results are shown in TABLE II below.

TABLE II

Sample	%C ^a CO	%C ^a HC	%C ^a SO ₂
S-3Ce	80.2	37.5	4.1
S-3	49.3	7.55	4.6
S-3Al	94.9	56.5	8.0

^a "%C" means the percentage conversion of the indicated constituent.

[0049] The data of TABLE II clearly indicate that the platinum is a more effective oxidation catalyst for HC and CO when dispersed on the alumina (S-3Al) as compared to when it is dispersed on the ceria (S-3Ce) and is much more effective in this regard than is the S-3 sample, wherein the platinum is dispersed equally on each of the ceria and alumina components. Overall, the best results were obtained with the S-3Ce sample in which fairly high levels of desired conversion of CO and HC were attained and the lowest level (4.1%) of the undesired oxidation of SO₂ to SO₃ was also attained. S-3 catalyst provided significant, but lesser, conversions of CO and HC and only slightly more (4.6%) of the undesired oxidation of SO₂ than did S-3Ce, but was much better in terms of less promotion of oxidation of SO₂ than was the S-3Al sample (8.0%). TABLE II thus demonstrates the desirability of dispersing all or at least a part of the platinum metal component on the ceria component of the ceria-alumina catalytic material.

Example 4

[0050] A series of catalyst samples was prepared generally according to the procedures of Example 1 to provide an alumina undercoat at a loading of 1.0 g/in³ (61 g/dm³) on which a metal oxide topcoat was coated. In the case of comparative sample Comp.1, the topcoat contained no ceria, the topcoat of comparative sample Comp.2 contained no alumina, and, in a third sample in accordance with the present invention, S-3, the topcoat comprised a ceria-alumina catalytic material containing 46.2 percent ceria and 53.8 weight percent alumina. Each of the samples contained 0.5 g/ft³ (18 g/m³) of platinum and had a topcoat loading of about 1.95 g/in³ (119 g/dm³), inclusive of the platinum. In all cases the ceria had a surface area of 164 m²/g and the alumina had a surface area of 150 m²/g. The samples were tested with the same test gas as described in Example 2 at 275°C, 350°C, 425°C and 500°C, and the conversion of HC, CO and oxidation of SO₂ to SO₃ at a space velocity of 50,000 was measured. The results of these tests are summarized in TABLE III.

TABLE III

Inlet Gas Temp. (°C)	Sample No.	%C SO ₂	%C HC	%C CO
275	Comp.1	16.3	10.0	96.6
275	S-3	0.0	2.4	30.5
275	Comp.2	10.2	0.0	9.4
350	Comp.1	18.9	86.5	99.6
350	S-3	4.0	9.8	68.3
350	Comp.2	12.2	6.5	63.1
425	Comp.1	35.5	90.5	99.9
425	S-3	11.8	31.6	84.3
425	Comp.2	22.4	18.2	70.4
500	Comp.1	42.2	83.7	99.7
500	S-3	12.0	47.4	84.8
500	Comp.2	32.0	31.6	61.0

[0051] The data of TABLE III indicate the conversion of hydro-carbons (HC) was highest for sample Comp.1, containing 100 percent alumina and no ceria, and lowest for sample Comp.2, containing 100 percent ceria and no alumina. The catalyst in accordance with the present invention, S-3, provided intermediate levels of conversion of HC. Comparable results were obtained for conversion of CO at all temperature levels. The results of TABLE III concerning the conversion of SO₂ to SO₃ are shown in the perspective-view plot of Figure 3 from which it is readily seen that at each temperature level tested a lower degree of conversion of SO₂ was attained by the S-3 sample in accordance with an embodiment of the present invention, than was attained with either the 100 percent alumina (Comp.1) version or the 100 percent ceria (Comp.2) version. These data demonstrate that utilizing a ceria-alumina catalytic material in accordance with the present invention reduces the oxidation of SO₂ as compared to either a 100 percent ceria or 100 percent alumina catalyst containing 0.5 g/ft³ (18 g/m³) of platinum.

[0052] A series of catalyst compositions was prepared in order to test catalyst compositions in accordance with the present invention against comparative catalyst compositions containing various refractory metal oxides and catalytic metals. These catalysts were tested both on a laboratory diagnostic reactor and on diesel engines. The two test engines employed were a Cummins 6BT engine, rated at 190 horsepower and having a 5.9 liter displacement and a Caterpillar 3176 engine, rated at 325 horsepower (242 kW) and having a 10.3 liter displacement. The operating characteristics of these two engines are shown in TABLE IV based on the operating cycle used to test the catalyst composition samples.

TABLE IV

Temperature (°C)	Caterpillar 3176 Temp. Cycle ^a		Cummins 6BT Temp. Cycle ^a	
less than 100	0		0	
100-200	0		62.6	
200-300	57.3		36.7	
300-400	30.9		0.7	
400-500	11.8		0	
Maximum Temperature (°C):	475		305	
	g/HP-hr ^b	Wt.% ^c	g/HP-hr ^b	Wt.% ^c
Particulates:				
VOF	0.036	21.6	0.066	38.4
Sulfate	0.005	3.1	0.003	2.0
Carbon/Other ^d	0.127	75.3	0.103	59.6
Totals	0.168	100.0	0.172	100.0
Gas Phase:				
HC	0.123	--	0.300	--
CO	3.48	--	1.50	--
NO _x	5.06	--	4.34	--

^a Percentage of cycle time at which the inlet exhaust to the catalyst lies within the indicated temperature range

^b "g/HP-hr" = grams per brake horsepower-hour of component emitted in exhaust (1 g/hp-hr = 0.373 g/MJ)

^c Weight percentage of total particulates provided by the indicated constituent

^d "Carbon/Other" values are calculated by difference between the measured VOF and sulfate components of the exhaust and the total exhaust particulates. Carbon/Other comprises the dry, solid carbonaceous content of the particulates plus any water associated with the sulfates. Any measurement errors will affect the "Carbon/Other" value.

[0053] As shown in TABLE IV, the Cummins engine runs with a cooler exhaust than does the Caterpillar engine and the total engine emissions are roughly comparable although the Cummins engine runs richer in the volatile organic fraction (VOF) which is the component most effectively treated by the diesel oxidation catalyst of the present invention.

Example 5

[0054] A series of catalyst samples was prepared generally by the method disclosed in Example 1 including two catalysts, designated samples S-3 and S-3B, comprising embodiments of the present invention and made exactly in accordance with Example 1 except that for sample S-3B palladium was substituted for platinum by using palladium nitrate as the source of the catalytic metal. Samples S-3 and S-3B each had an alumina undercoat at a loading of 1.0 g/in³ (61 g/dm³) and a topcoat of the ceria-alumina coating at a loading of 1.95 g/in³ (119 g/dm³). A series of comparative catalysts designated Comp.4, Comp.4M, Comp.4B, Comp.7, Comp.2.3, Comp.6 and Comp.5 were made by procedures comparable to those used in Example 1, with the following differences. The comparative catalysts were made without an alumina undercoat and, of course, using different refractory metal oxides as indicated by their respective compositions. For the samples containing niobia-silica (Comp.4, 4M, 4B and 7) the niobia was provided by dissolving niobium oxalate in the coating slurry. Further, the foamed α -alumina ("FAA") of Comp.2.3 and the silica of other comparative samples were not ball milled but were dry-jet milled and then incorporated into the coating step by use of a high speed intensive mixer. The vanadia-titania of sample Comp.6 was incorporated into a slurry containing palladium nitrate as the catalytic metal source.

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[0055] The silica employed in each case except Comp.2.3 was an extremely porous silica designated PQ-1022 by its manufacturer, PQ Corporation. The PQ-1022 silica has a porosity of 1.26 cc/g pore volume comprised of pores having a radius of from about 10 to 300 Angstroms (1 to 30 nm), and a surface area of 225 m²/g. The high porosity of the silica accounts for the relatively low weight loadings of the silica-containing washcoats. A silica sol was used for the Comp. 2.3 sample as described in footnote c of TABLE V. Each of these catalysts, the general composition of which is set forth in TABLE V, was prepared as a slurry of the refractory metal oxide or oxides indicated in TABLE V which had been impregnated with the specified loading of catalytic metal and then coated onto 400 cpsi (6200 cells/dm²) cylindrical cordierite honeycomb carriers manufactured by NGK and measuring 9 inches (22.86 cm) in diameter by 6 inches (15.24 cm) in length, providing a catalyst volume of 6.25 liters.

TABLE V

Catalyst Sample	Washcoat	Metal	Metal Loading (g/ft ³) ^f	Hours Aged	
				24	100
S-3	Ceria-Alumina	Pt	0.5	X	X
Comp.4	Niobia-Silica ^a	Pd	50.0	X	X
Comp.4M	Niobia-Silica ^a	Pd-Pt	25-5	X	X
S-3B	Ceria-Alumina	Pd	50.0	X	X
Comp.4B	Niobia-Silica ^a	Pt	0.5	X	
Comp.7	MnO-Niobia-Silica ^b	Pt	2.2	X	
Comp.2.3 ^c	Silica-FAA ^d	Pt	2.2	X	
Comp.6	Vanadia-Titania ^e	Pd	27	X	X

^a The niobia-silica sample catalysts (Comp.4, 4M and 4B) had washcoats comprised of 10 percent by weight niobia and 90 percent by weight silica, with a total washcoat loading of 0.8 g/in³ (49 g/dm³).

^b The MnO-niobia-silica sample catalyst (Comp.7) had a washcoat comprised of 90 percent by weight silica, 4 percent by weight niobia and 6 percent by weight MnO, with a total washcoat loading of 0.6 g/in³ (37 g/dm³).

^c The silica-foamed α -alumina sample catalyst (Comp.2.3) had a washcoat comprised of 10 percent by weight silica sol binder and 90 percent by weight of foamed α -alumina ("FAA"), with a total washcoat loading of 0.6 g/in³ (37 g/dm³). The α -alumina has a porosity of 0.0439 cc/g pore volume comprised of pores having a radius of from about 10 to 300 Angstroms (1 to 30 nm), and a surface area of 20.3 m²/g.

^d "FAA" = foamed α -alumina

^e The vanadia-titania sample catalyst (Comp.6) had a washcoat comprised of 4 percent by weight vanadia and 96 percent by weight titania, with a total washcoat loading of 1.8 g/in³ (110 g/dm³).

^f 1 g/ft³ = 35.315 g/m³

[0056] All eight sample catalysts were evaluated on the Cummins 6BT engine employing the U.S. Transient Cycle (commonly, and sometimes hereinbelow, referred to as the "Federal Test Procedure" or "FTP"). A description of the U.S. Transient Cycle is set forth in the Code of Federal Regulations, Title 40, Chapter 1, Subpart N, Paragraphs 86:1310-88 and 86:1312-88, Appendix I(f)(2). The catalyst volume-to-engine displacement ratio was 1.06. The catalysts were evaluated for fresh activity (after 24 hours aging) following which the five indicated samples were aged for 100 hours and further evaluated. All catalysts were aged on a 1986 Cummins NTC diesel engine rated at 400 horsepower (298 kW) and having a 14.0 liter displacement. The aging cycle employed flowed the engine exhaust through three catalysts of 6.25 liter volume each, simultaneously and in parallel, with the engine load adjusted to provide fifteen minute cycles during which the exhaust attained inlet temperatures as follows for the indicated amount of time:

330 - 400°C for 14% of the time,
 400 - 500°C for 22% of the time,
 500 - 550°C for 50% of the time, and
 550 - 565°C for 14% of the time.

[0057] The S-3 and S-3B samples each contain 46.2 weight percent aluminum-stabilized ceria and 53.8 weight percent alumina.

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[0058] TABLE VI shows the results of the fresh (aged 24 hours) catalyst samples tested under the Federal Test Procedure on the Cummins 6BT engine with all recorded exhaust emissions being given in grams per brake horsepower-hour (1 g/hp-hr = 0.373 g/MJ). All emissions are measured quantities except for "Carbon + Other" which is calculated by difference. The measured values are the average of four different runs conducted under the Federal Test Procedure which were carried out over the space of two days in order to account for day-to-day variations. TABLE VI also shows the base line values of the diesel exhaust operated without catalytic treatment over an average of 24 runs. The difference between the runs carried out without catalytic treatment and the runs carried out using the various catalyst samples were utilized to calculate the percent conversion of each of the emissions components. The percent conversion is the percentage of the emissions contained in the untreated exhaust which were converted to innocuous components by utilization of the catalyst samples. The abbreviation "TPM" is used for "total particulate matter".

TABLE VI

Catalyst Sample	HC	CO	NOx	TPM	VOF	Sulfate	Carbon + Other
None - Untreated engine exhaust							
Grams ^a	0.299	1.5	4.34	0.172	0.0611	0.0034	0.108
S-3							
Grams ^a	0.188	1.11	4.3	0.118	0.0256	0.0016	0.0908
%C ^b	37.4	26	0.96	31.7	58.1	53.1	15.9
Comp.4							
Grams	0.198	1.28	4.22	0.123	0.0272	0.0022	0.0936
%C	34.1	14.9	2.7	28.8	55.4	37	13.3
Comp.4M							
Grams	0.213	1.34	4.22	0.123	0.0302	0.0025	0.0903
%C	29.1	11	2.8	28.8	50.6	28.2	16.4
S-3B							
Grams	0.155	1.31	4.31	0.118	0.0258	0.0025	0.0897
%C	48.3	13	0.73	31.7	57.7	26	16.9
Comp.4B							
Grams	0.208	1.17	4.27	0.128	0.0359	0.0033	0.0888
%C	30.7	22.2	1.5	25.9	41.2	4.8	17.8
Comp.7							
Grams	0.198	1.09	4.31	0.135	0.0378	0.0028	0.0944
%C	34.1	27.2	0.79	21.5	38.1	17.2	12.6
Comp.2.3							
Grams	0.185	1.09	4.34	0.135	0.0306	0.0049	0.0995
%C	38.2	27.5	0.1	21.5	38.1	-43.6	7.9
Comp.6							
Grams	0.135	1.51	4.35	0.118	0.0255	0.003	0.0895
%C	54.9	-0.3	-0.25	31.7	58.3	11.4	17.1

^a Grams per brake horsepower-hour (1 g/hp-hr = 0.373 g/MJ)

^b "%C" means the percentage conversion of the indicated constituent. A negative %C means the treated exhaust contained more of the constituent than did the untreated exhaust.

[0059] The results tabulated in TABLE VI indicate that with respect to VOF conversion and total particulates conver-

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sion, the best results were obtained by S-3, S-3B and Comp.6 catalysts, with the Comp.4 sample giving the next best results. As to sulfate emissions, the Comp.2.3 sample exhibited sulfate emissions which were greater than those of the untreated exhaust, all the other samples tested giving at least some reduction in sulfates as compared to the untreated exhaust. This finding is consistent with the relatively low temperature of the Cummins 6BT engine. With respect to gas phase emissions (HC, CO and NO_x) Comp.6, S-3B and Comp.5 gave the best HC reduction while Comp.2.3, Comp.7 and S-3 gave the best CO conversion. There was little catalytic effect on NO_x emissions as one would expect in the relatively oxygen-rich environment of a diesel exhaust.

Example 6

[0060] As indicated in TABLE V, five of the catalysts tested were then aged to a total of 100 hours and re-evaluated on the Cummins 6BT engine. The results of the evaluation of the 100-hour aged samples are summarized in TABLE VII.

TABLE VII

Catalyst Sample	HC	CO	NO _x	TPM	VOF	Sulfate	Carbon + Other
None - Untreated engine exhaust							
Grams ^a	0.305	1.55	4.46	0.179	0.0675	0.0039	0.108
S-3							
Grams ^a	0.188	1.27	4.31	0.123	0.0284	0.0018	0.0928
%C ^b	38.4	17.9	3.3	31.3	57.9	53.8	14.1
Comp.4							
Grams	0.218	1.47	4.37	0.128	0.0327	0.0023	0.093
%C	28.5	4.9	1.9	28.5	51.6	41	13.9
Comp.4M							
Grams	0.238	1.49	4.37	0.13	0.0349	0.0031	0.092
%C	22	3.6	1.9	27.4	48.3	20.5	14.8
S-3B							
Grams	0.175	1.27	4.38	0.12	0.0282	0.0022	0.0896
%C	42.6	17.9	1.7	33	58.2	43.6	17
Comp.6							
Grams	0.22	1.69	4.42	0.14	0.0308	0.0042	0.105
%C	27.9	-9.3	0.8	21.6	54.4	-7.7	2.8

^a Grams per brake horsepower-hour (1 g/hp-hr = 0.373 g/MJ)

^b "%C" means the percentage conversion of the indicated constituent. A negative %C means the treated exhaust contained more of the constituent than did the untreated exhaust.

[0061] Table VII shows that the best results were attained by the S-3 and S-3B catalysts for both total particulate emissions and VOF conversion. With respect to HC reduction the best performance was shown by S-3B although the S-3 catalyst proved to be the most stable, the results attained by the S-3 catalyst after 100 hours aging being actually better than those attained by the 24-hour aged S-3 sample. The S-3B catalyst exhibited improved CO conversion for the 100-hour aged catalyst as compared to the fresh (24-hour aged) catalyst. Note that the Comp.6 sample removed essentially no CO at 24 hours and became a net CO producer after being aged for 100 hours. The results of TABLE VI and VII clearly show that the catalyst compositions of the present invention, S-3 and S-3B, gave the best overall emissions control and the best durability as evidenced by 100 hours of aging.

Example 7

[0062] In order to compare the effect of different catalytic metal loadings on the performance of catalysts in accord-

ance with the present invention, three sample catalysts in accordance with the present invention were prepared in accordance with the procedure of Example 1. Thus, each catalyst comprised a cordierite 400 cpsi (6200 cells/dm²) substrate containing 1.95 g/in³ (119 g/dm³) of the ceria-alumina catalytic material of the invention. The ceria-alumina catalytic material contained 46.2 weight percent of aluminum-stabilized ceria and 53.8 weight percent of activated alumina. Each catalyst had an alumina undercoat in the amount of 1.00 g/ft³ (35 g/m³) onto which the ceria-alumina catalytic material was coated. One sample, designated S-3.5Pt had 0.5 g/ft³ (35 g/m³) of platinum dispersed thereon, another sample, designated S-3.20Pt had 2.0 g/ft³ (71 g/m³) of platinum dispersed thereon and a third sample, designated S-3Pd had 50 g/ft³ (1766 g/m³) of palladium dispersed thereon. Each catalyst was tested under the Federal Test Procedure to treat an exhaust generated by a Cummins C-series 250 HP diesel engine having a displacement of 8 liters, so that a catalyst volume-to-engine displacement ratio of 0.78 was utilized. The effectiveness of the sample catalyst was tested in the same manner as that of Example 6 and the results with respect to conversion of total particulates (TPM) and gaseous phase HC and CO are set forth in TABLE VIII.

TABLE VIII

Sample	%C ^a TPM	%C ^a HC	%C ^a CO
S-3.5Pt	47	28	7.5
S-3.20Pt	48	69	74
S-3Pd	48	52	35

^a "%C" means the percentage conversion of the indicated constituent.

[0063] The data of TABLE VIII show that all three samples were nearly identical with respect to the percentage conversion of total particulates although the larger loadings of catalytic metal made a dramatic difference in the percentage conversions of the gaseous HC and CO. These results are consistent with the data of Example 6 and TABLE VII, from which it will be noted that S-3 and S-3B gave substantially similar results with respect to total particulates reduction in spite of the fact that S-3 contains only 0.5 g/ft³ (18 g/m³) of platinum and S-3B contains 50 g/ft³ (1766 g/m³) of palladium. The lack of pronounced effect on total particulate reduction between a catalyst containing 100 times more platinum group metal than another, strongly suggests the irrelevancy of the presence of the catalytic metal insofar as total particulate reduction is concerned, and that particulate reduction is attained by the effect of ceria-alumina catalytic material.

Example 8

[0064] In order to further demonstrate the irrelevancy of the platinum metal loading insofar as catalytic activity of the ceria-alumina catalytic material with respect to total particulate reduction is concerned, a series of samples of catalytic material powder was prepared. This was done by utilizing the ceria-alumina washcoat material of Example 7 containing various quantities of platinum metal ranging from 0 to the equivalent of 5.0 g/ft³ (177 g/m³) of platinum if the washcoat were to be coated upon a 400 cpsi (6200 cells/dm²), NGK cordierite substrate. The resultant series of powders were each mixed with 10 weight percent of a diesel engine lubricating oil, Cummins SAE-15W Premium Blue Diesel Engine Lube Oil, and the sample of the mixture was evaluated by simultaneous thermogravimetric analysis and differential thermal analysis (TGA/DTA) for combustion of the lubricating oil. It should be noted that unburned diesel engine lubricating oil constitutes a significant portion of the volatile organic fraction (VOF) of diesel exhaust particulate emissions and the efficacy of the ceria-alumina catalytic material in catalyzing combustion of the lubricating oil is a good indication of the effectiveness of the ceria-alumina catalytic material in catalyzing oxidation of VOF and, thereby, reduction of particulate emissions. Thermogravimetric analysis measures the weight gain or loss of a sample (indicating a chemical reaction undergone by the sample) as a function of the temperature to which the sample is heated. Differential thermal analysis measures the amount of energy (heat) absorbed by the sample (indicating that the sample has undergone an endothermic reaction) or liberated by the sample (indicating that the sample has undergone an exothermic reaction) as a function of the temperature to which the sample is heated. Figure 4 is a plot of the results obtained by heating the mixture of catalytic material powder and lubricating oil in a temperature regime ranging from ambient temperature to 600°C and recording the TGA/DTA data. The DTA peak area was corrected for the weight change determined by the TGA so that the results attained are proportional to the amount of lubricating oil combusted, i.e., to the effectiveness of the tested ceria-alumina catalytic materials, which are identical except for the varying platinum metal loadings. The results attained are plotted in Figure 4 wherein, despite some experimental scatter in the data points, the trend line indicates

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substantially no effect of the platinum content of the catalytic material insofar as lubricating oil combustion is concerned. Thus, about the same proportion of combustion was attained for the ceria-alumina catalytic material containing no platinum as for that containing incremental amounts of platinum up to and including the equivalent of 5 g/ft³ (177 g/m³) on a 400 cpsi (6200 cells/dm²) carrier.

Example 8A

[0065] An equivalent test of silica based and silica-niobia based refractory metal oxide powders on which varying amounts of platinum were dispersed was carried out. Those test results showed that the ceria-alumina catalytic material of the present invention provided better performance for lubricating oil combustion as measured by DTA and therefore, by implication, for catalytic oxidation of VOF in diesel engine exhaust.

Example 9

[0066] S-3 and comparative Comp. 4 catalyst samples were tested on the exhaust of the Caterpillar 3176 engine. As previously noted, this engine runs with a considerably hotter exhaust than the Cummins 6BT engine and test catalysts of the same size (9 inches (22.86 cm) by 6 inches (15.24 cm) providing a catalyst volume of 6.25 liters) were tested on this larger engine, providing a catalyst volume-to-engine displacement ratio of 0.607. S-3 and Comp.4 catalyst samples were aged for 50 hours on an aging cycle similar to that described in Example 5 but from a minimum of about 300°C to a maximum of about 530°C.

[0067] The results of this test are shown in TABLE IX as the average of six hot-start runs in accordance with the Federal Test Procedure.

TABLE IX

Catalyst Sample	HC	CO	NOx	TPM	VOF	Sulfate	Carbon + Other
None - Untreated engine exhaust							
Grams ^a	0.123	3.48	5.06	0.168	0.0363	0.0052	0.1265
S-3							
Grams ^a	0.1566	2.5	4.95	0.138	0.0213	0.0039	0.1128
%C ^b	54	28.2	2.2	17.9	41.3	25	10.8
Comp.4							
Grams	0.0833	1.76	5.02	0.177	0.017	0.0217	0.1383
%C	32.3	20.7	0.8	-5.4	53.2	-317	-9.3

^a Grams per brake horsepower-hour (1 g/hp-hr = 0.373 g/m³)

^b "%C" means the percentage conversion of the indicated constituents. A negative %C means the treated exhaust contained more of the constituent than did the untreated exhaust.

[0068] The results summarized in TABLE IX show that the S-3 catalyst reduced total particulate emissions by 17.9 percent and VOF by 41.3 percent whereas the Comp.4 sample, although it gave a higher VOF reduction at 53.2 percent, resulted in an increase of total particulate emissions, because of its very high sulfate make which resulted in sulfate emissions 317 percent higher than those emitted in the untreated exhaust. The tendency of the Comp.4 sample to produce large amounts of sulfate in the hot exhaust environment of the Caterpillar 3176 engine stands in marked contrast to the efficiency of the S-3 catalyst in attaining a 25 percent reduction in sulfate emissions and therefore an overall reduction in total particulates. The fact that the S-3 catalyst exhibited lower total particulates and VOF removal levels on the Caterpillar 3176 engine than on the Cummins 6BT engine is attributable to the smaller catalyst volume relative to engine size encountered on the Caterpillar engine test and to the fact that the concentration of VOF, the component most vigorously treated by the catalyst, is some 40 percent lower in the exhaust of the Caterpillar engine than in the exhaust of the Cummins engine.

Example 10

[0069] In order to compare the effect on conversion of SO₂ to SO₃, and thus sulfate-make of a catalyst, three com-

parative samples, one of which (designated Comp.11) is a commercially available diesel exhaust catalyst, and each containing high platinum group metal loadings, were compared to a fourth sample comprising an embodiment of the present invention. Three samples, Comp.1, Comp.2 and S-3, were prepared generally in accordance with the procedure of Example 1 to coat cylindrical carriers comprising 400 cpsi (6200 cells/dm²) cordierite cores measuring 1.5 inches (3.81 cm) in diameter by 3 inches (7.62 cm) in length. The samples were aged for 10 hours at 500°C by having a mixture of 10 percent steam in air flowed through each sample. Comparative sample Comp.1 comprised 50 g/ft³ (1766 g/m³) of platinum disposed on an activated alumina carrier and comparative sample Comp.2 had a 50 g/ft³ (1766 g/m³) platinum group metal loading, the platinum group metal comprising platinum and rhodium in a 5:1 weight ratio disposed on a ceria-alumina catalytic material comprising 53.8 percent by weight alumina and 46.2 percent by weight aluminum-stabilized ceria. The S-3 sample comprised, in accordance with one embodiment of the present invention, 0.5 g/ft³ (18 g/m³) of platinum dispersed on a ceria-alumina catalytic material comprising 46.2 percent by weight aluminum-stabilized ceria and 53.8 percent by weight alumina with one-half the platinum metal dispersed on the aluminum-stabilized ceria and one-half the platinum metal dispersed on the alumina. The commercially available catalyst for diesel exhaust applications was analyzed and found to comprise a catalytic material dispersed on a honeycomb-type carrier having 400 cells per square inch (6200 cells per dm²). The commercial catalyst contained about 50 g/ft³ (1766 g/m³) of platinum dispersed on a support comprised primarily of titania, vanadia and alumina. A core 2.5 inches (6.32 cm) long and 1.5 inches (3.81 cm) in diameter was cut from the commercial catalyst and this comparative catalyst core was designated as Comp.11. The four catalyst samples were tested at space velocities of 50,000 and 90,000 at temperatures of 275°C, 350°C, 425°C and 500°C. (In this Example 10 and in Example 11 below, the flow rate of the reaction gas was adjusted as necessary to compensate for the slight difference in catalyst volume so that each tested sample was evaluated at the same space velocity.) Each sample was held at the indicated temperature for 10 minutes during the evaluation. The test gas used in the laboratory diagnostic unit comprised 10 percent steam, 10 percent oxygen, 4.5 percent CO₂, 1000 ppm NO, 28.57 ppm heptane (equivalent to 200 ppm C₇ hydrocarbons), 28.6 ppm CO, 50 ppm SO₂, balance nitrogen. (The percents are volume percents and "ppm" means parts per million by volume.) The results of these evaluations are given in TABLE X.

TABLE X

Catalyst Sample/Inlet Temp.	Percent Conversion at Indicated Space Velocity					
	50,000 SV			90,000 SV		
	CO	HC	SO ₂	CO	HC	SO ₂
Comp. 1						
275°C	99.5	68.6	56.9	94.9	52.4	29.4
350°C	99.5	83.0	76.9	96.5	70.5	58.5
425°C	100	87.4	94.3	--	--	--
500°C	100	89.0	92.2	--	--	--
Comp.2						
275°C	100	52.5	--	98.5	47.6	8.0
350°C	99.0	77.5	11.8	96.9	61.7	9.8
425°C	99.0	84.2	31.4	96.0	74.4	23.5
500°C	98.1	90.7	47.1	95.5	73.2	37.3
Comp.11						
275°C	97.1	16.7	0.0	84.7	4.8	2.0
350°C	99.0	54.5	2.0	93.0	41.0	2.0
425°C	99.5	75.0	23.6	97.0	63.2	15.7
500°C	99.5	85.9	54.9	97.4	73.2	38.0
S-3						
275°C	30.5	2.4	0.0	10.9	0.0	0.0
350°C	68.3	9.8	4.0	52.4	9.5	0.0
425°C	84.3	31.6	11.8	59.5	24.3	3.9
500°C	84.4	47.4	12.0	60.0	28.6	4.1

[0070] The data of TABLE X shows that the comparative samples Comp.1 and Comp.2 exhibit very high conversion of SO₂ to SO₃, and thus high sulfate make, even at the lowest test temperature of 275°C and high space velocity of 90,000. Although comparative sample Comp.2 exhibits less sulfate make than Comp.1 (but significantly more than catalyst S-3, discussed below), this is believed to be due primarily to the modifying effect of rhodium on the SO₂ oxidation activity of platinum. The Comp.2 catalyst has the economic disadvantage of being too costly because of the very high cost of rhodium even as compared to the cost of platinum. Both comparative samples Comp.1 and Comp.2 show high HC and CO conversion. S-3, the sample in accordance with an embodiment of the present invention, exhibits greatly reduced SO₂ conversion relative to both Comp.1 and Comp.2 with practically no SO₂ conversion occurring in the low temperature regime and with relatively small SO₂ conversion even at the high temperature regime. Some activity for conversion of gaseous HC and CO is exhibited by catalyst S-3, especially at the high temperature regime where good CO and moderate HC activity is seen. The data of TABLE X thus clearly demonstrate that the utilization of a low platinum loading on the ceria-alumina catalytic material of the present invention provides excellent control of SO₂ oxidation and consequently excellent control of total particulates emission in a diesel engine exhaust. It should be noted that the diagnostic test is a very stringent test of sulfate oxidation as compared to actual engine performance. Experience has shown that a given catalyst will perform better with respect to sulfate oxidation in treating an actual diesel engine exhaust than it will in the diagnostic engine test.

[0071] The comparative catalyst sample Comp.11 is seen to suppress SO₂ oxidation in a manner comparable to that of sample S-3, but only up to a temperature between 350° and 425°C. At 425°C and higher temperatures the Comp.11 sample exhibits greatly increased SO₂ oxidation as compared to the S-3 catalyst sample. Accordingly, the catalyst sam-

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ple of the present invention, even with a 0.5 g/ft^3 (18 g/m^3) platinum loading, appears to be significantly better with respect to SO_2 oxidation in higher temperature regimes than the commercial catalyst of Comp.11. The comparative samples Comp.1, Comp.2 and Comp.11 all contain high platinum loadings and consequently show higher HC and CO conversion than does the 0.5 g/ft^3 (18 g/m^3) platinum S-3 catalyst sample. However, as pointed out elsewhere herein, HC and CO emissions generally satisfactorily controlled by engine design and the problem which the art is seeking to overcome is to control the total particulates emissions which, as noted above, is in part a function of sulfate make. The catalysts of the present invention, with no or a very low loading of platinum, show excellent activity for reducing total particulate emissions because of their unexpected good activity for oxidizing VOF and their low sulfate make. Further, it is obviously economically advantageous to eliminate or drastically reduce the platinum metal loading in accordance with the teachings of the present invention.

Example 11

[0072] A catalyst sample in accordance with an embodiment of the present invention was prepared and designated sample S-3P. Catalyst sample S-3P is identical to catalyst sample S-3 of Example 10 except that it contains a platinum loading of 2.0 g/ft^3 (71 g/m^3). The S-3P catalyst sample was 3.0 inches (7.62 cm) long by 1.5 inches (3.81 cm) in diameter. Catalyst S-3P was tested by passing therethrough a test gas in the same manner as described in Example 10 at space velocities of 50,000 and 90,000 at temperatures of 275°C , 350°C , 425°C and 500°C . The results of this test are shown in TABLE XI.

TABLE XI

Catalyst Sample/Inlet Temp.	Percent Conversion at Indicated Space Velocity					
	50,000 SV			90,000 SV		
	CO	HC	SO_2	CO	HC	SO_2
S-3P						
275°C	99.0	10.0	16.0	88.5	2.6	2.0
350°C	100	87.8	21.6	98.0	74.1	5.9
425°C	100	90.5	33.3	98.5	82.7	19.2
500°C	99.5	83.1	35.3	98.1	73.3	31.4

[0073] TABLE XI shows, as would be expected, that the S-3P catalyst exhibits higher SO_2 oxidation at all temperature levels and space velocities as compared to the S-3 catalyst of Example 10 which contains 0.5 g/ft^3 (18 g/m^3) of platinum, one-fourth of the amount of platinum (2.0 g/ft^3 (71 g/m^3)) which S-3P contains. However, the S-3P sample also exhibited higher HC and CO conversions, which shows that a modest increase in platinum loading, still keeping the total platinum loading to very low levels as compared to prior art catalysts, can accommodate a higher HC and CO conversion but at the potential cost of somewhat increased particulate emissions because of additional sulfate make. However, in certain circumstances it may be desirable to attain the higher HC and CO conversions attainable with the catalyst of the present invention by a modest increase in platinum loading.

Example 12

[0074] In order to evaluate the effect of ceria in the catalyst composition of the present invention, a comparative sample, Comp.1 of Example 4, was prepared generally in accordance with the procedure of Example 1 but omitting the ceria component of the catalytic material. Thus, the resulting catalyst comprised an activated alumina washcoat having 0.5 g/ft^3 (18 g/m^3) of platinum disposed thereon. This sample designated Comp.3C was subjected to the same test as in Examples 10 and 11 and the results thereof are summarized in TABLE XII and show that the SO_2 conversion over this catalyst is significantly greater than for the S-3 catalyst, especially at low temperatures. Higher conversions of HC and CO were also attained. This data clearly indicate that the ceria plays an important modifying role in the oxidation activity of the platinum.

TABLE XII

Catalyst Sample/Inlet Temp.	Percent Conversion at Indicated Space Velocity					
	50,000 SV			90,000 SV		
	CO	HC	SO ₂	CO	HC	SO ₂
Comp.3C						
275°C	96.6	10.0	16.3	85.4	4.9	4.6
350°C	99.6	86.5	18.9	95.0	57.4	12.6
425°C	99.9	90.5	35.5	98.1	74.0	33.7
500°C	99.7	83.7	42.2	98.3	74.0	33.7

Example 13

[0075] A catalyst was prepared in accordance with the present invention generally following the teachings of Example 1, except that no alumina undercoat was utilized. Thus, this sample comprised 1.95 g/in³ (119 g/dm³) of a ceria-alumina catalytic material containing 46.2 weight percent aluminum-stabilized ceria (164 m²/g BET surface area) and 53.8 weight percent alumina (150 m²/g BET surface area) disposed directly upon the carrier without an alumina undercoat, and having 0.5 g/ft³ (18 g/m³) of platinum dispersed thereon. This catalyst, designated S-3SC was aged and tested in the same manner as in Example 10 and the results thereof are shown in TABLE XIII. The performance of this sample is seen to be essentially the same as that of S-3 (Example 10, TABLE X) for the gas phase reactions, indicating that the presence of the alumina undercoat is not essential with respect to either low sulfate make or HC and CO oxidation.

TABLE XIII

Catalyst Sample/Inlet Temp.	Percent Conversion at Indicated Space Velocity					
	50,000 SV			90,000 SV		
	CO	HC	SO ₂	CO	HC	SO ₂
S-3SC						
275°C	25.4	0.0	2.0	31.0	0.0	0.0
350°C	71.9	11.9	5.9	62.5	15.8	4.1
425°C	85.6	28.9	9.8	78.7	29.3	5.9
500°C	86.3	48.7	20.4	76.1	42.5	10.7

Example 14

[0076] A. Catalysts were prepared generally in accordance with the procedures of Example 1 to provide a series of three otherwise identical compositions containing a ceria-alumina catalytic material in accordance with the teachings of the present invention having platinum dispersed thereon, including 0.0, 0.5 and 2.0 g/ft³ (0, 18 and 71 g/m³) platinum. Each catalyst comprised a γ -alumina undercoat at a loading of 1.0 g/in³ (61 g/dm³) upon which was coated a top layer comprised of 1.05 g/in³ (64 g/dm³) γ -alumina plus 0.90 g/in³ (55 g/dm³) alumina-stabilized ceria (2.5 weight percent Al₂O₃ based on the combined weight of bulk ceria and alumina dispersed therein). The catalysts were coated onto a 9 inch (22.86 cm) diameter by 6 inch (15.24 cm) long, 400 cpsi (6200 cells/dm²) cordierite substrate. The resulting catalyst samples were designated as S-4 (0.0 g/ft³ platinum, aged 24 hours), S-5 (0.5 g/ft³ (18 g/m³) platinum, aged 25 hours) and S-6 (2.0 g/ft³ (71 g/m³) platinum, aged 24 hours).

[0077] B. The three catalyst samples were conditioned prior to evaluation using an aging cycle involving 20 minutes each at Modes 2, 6 and 8 of the European 13 Mode Test Procedure (ECE R.49 Thirteen Mode Cycle). This Test Procedure is set forth in the Society of Automotive Engineers Publication, SAE Paper #880715, published at the International

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Congress and Exposition, Detroit, Michigan, February 29 through March 4, 1988, by Giorgio M. Cornetti et al. Prior to testing to develop the data of TABLE XV and Figures 5-8, the three catalyst samples were aged 24 or 25 hours as indicated below on a Cummins 6BT turbocharged diesel engine having a 5.9 liter displacement and rated at 190 horsepower (142 kW). For both aging and test purposes, the engine was run with low sulfur fuel (0.05 weight percent sulfur) under steady state conditions using test modes selected from the aforesaid European 13 Mode Cycle Test Procedure.

[0078] The engine conditions for the test modes along with average (for five runs) catalyst inlet temperatures and baseline emissions (of untreated engine exhaust) are shown in TABLE XIV.

TABLE XIV

Cummins 6BT 190 HP (142 kW) Turbocharged Diesel Engine, 5.9 Liter Displacement, Conditions For Steady State Catalyst Tests Engine Conditions				
Test Mode No.	rpm	% Load	Average Catalyst Inlet Temp. (°C)	
8	2515	100	571±2	
10	2515	50	338±4	
6	1609	100	549±5	
4	1609	50	400±4	
2	1560	10	214±3	
1	803	Low	128±16	
Baseline Emissions - Untreated Exhaust				
Test Mode No.	Average Emissions (g/bhp-hr) ¹			
	TPM	SOF	HC	CO
8	0.097	0.010	0.122	0.46
10	0.151	0.047	0.212	0.68
6	0.221	0.016	0.099	2.23
4	0.146	0.023	0.103	0.52
2	0.265	0.137	0.541	2.57
1	---	0.078	1.04	3.01

¹ grams per brake horsepower hour (1 g/hp-hr = 0.373 g/MJ)

[0079] The conditioned and aged catalyst samples S-4, S-5 and S-6 were tested for conversion of emission components in diesel exhaust generated by the test engine used to generate the data of Table XIV, as a function of steady state test mode and catalyst inlet temperature, i.e., the temperature of the diesel engine exhaust introduced to the catalyst. The results are summarized in TABLE XV.

TABLE XV

Sample/(Pt Load g/ft ³) ^a	Test Mode No.	Cat. Inlet Temp. °C	% Removal			
			SOF	TPM	HC	CO
S-4 (0.0)	2	209	72	63	31	1
	10	335	60	27	32	7
	4	399	62	18	38	18
	6	547	84	-40	44	27
	8	572	79	-181	39	-4
S-5 (0.5)	2	215	60	45	27	6
	10	343	58	28	41	63
	6	549	91	-64	56	85
	8	570	80	-201	62	45
S-6 (2.0)	1	127	56	52	37	-1
	2	215	61	61	39	8
	10	341	53	31	74	86
	4	397	61	22	82	87
	6	554	89	-60	78	95
	8	572	79	-200	71	70

^a 1 g/ft³ = 35.315 g/m³

[0080] The data of TABLE XV show that all three catalysts are comparable in SOF removal performance as a function of temperature, with the catalyst containing no platinum (S-4) performing as well as the catalysts containing platinum (S-5 and S-6).

[0081] With reference to TABLE XV and Figures 5-8, it is seen that the SOF removal performance as a function of inlet temperature of the three catalysts are compared in Figure 5. As can be seen, all three samples are comparable across the temperature range of about 120 to 575°C with the sample containing no platinum (S-4) performing as well as or better than, the platinum-containing samples S-5 and S-6.

[0082] The SOF removal performance is also reflected in the total particulate (TPM) removal levels of the three catalysts which are compared in Figure 6. The platinum-free catalyst sample is comparable to, or better than, the platinum-containing catalyst samples S-5 and S-6 at all temperatures. Note also, all three catalyst samples make particulates at the highest temperatures of the test. This is due to sulfate-make from the oxidation of gas phase SO₂ to SO₃. Thus, even the platinum-free sample makes sulfate at extremely high temperatures, but apparently to a slightly lesser extent than the platinum-containing samples, reflecting the lower gas phase activity of the platinum-free sample.

[0083] The gas phase activity of the three catalyst samples are compared in Figures 7 and 8 for, respectively, hydrocarbon ("HC") and carbon monoxide ("CO") gas phase conversions. Although the platinum-free sample S-4 exhibits some gas phase activity for HC and CO conversion, it is clear from these results that the platinum-containing samples S-5 and S-6 have substantially higher gas phase activity. This is especially clear in the case of CO conversion. The platinum-free sample S-4 has some gas phase activity because the ceria component has activity as an oxidation catalyst.

[0084] These results show quite well the surprising finding that a platinum-free catalyst in accordance with the present invention exhibits very good particulates removal from diesel engine exhaust because of its activity for the removal and combustion of VOF, and that the precious metal is not needed to accomplish this function. If there is a need to enhance gas phase HC and CO activity, this can be accomplished separately by adding a limited amount of platinum to the catalyst.

Claims

1. An oxidation catalyst composition that comprises a refractory carrier on which is disposed a coating of a ceria-alumina catalytic material consisting essentially of a combination of bulk ceria having a BET surface area of at least

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10 m²/g and bulk alumina having a BET surface area of at least 10 m²/g.

2. The catalyst composition of claim 1 wherein the ceria and alumina each comprises from 5 to 95 percent by weight of the combination.
- 5 3. The catalyst composition of claim 1 wherein the ceria and alumina each comprises from 10 to 90 percent by weight of the combination.
- 10 4. The catalyst composition of claim 1 wherein the ceria and alumina each comprises from 40 to 60 percent by weight of the combination.
5. The catalyst composition of any of claims 1 to 4 wherein the ceria and the alumina are each disposed in respective discrete layers, one overlying the other.
- 15 6. The catalyst composition of any of claims 1 to 5 wherein the ceria comprises an aluminum-stabilized ceria.
7. The catalyst composition of any of claims 1 to 6 wherein the ceria and the alumina each has a BET surface area of from 25 m²/g to 200 m²/g.
- 20 8. The catalyst composition of any of claims 1 to 7 further comprising catalytically active platinum or palladium dispersed on the catalytic material, the platinum or palladium being present in a moderate amount to limit or to reduce the oxidation of SO₂ to SO₃.
- 25 9. The catalyst composition of claim 8 wherein the catalytic material has dispersed thereon a catalytically effective amount of platinum not exceeding 25 g/ft³ (883 g/m³) of the catalyst composition.
10. The catalyst composition of claim 8 wherein the catalytic material has dispersed thereon a catalytically effective amount of platinum not exceeding 15 g/ft³ (530 g/m³) of the catalyst composition.
- 30 11. The catalyst composition of claim 8 further comprising platinum dispersed on the catalytic material, the platinum being present in the amount of from 0.1 to 15 g/ft³ (3.5 to 530 g/m³) of the composition.
12. The catalyst composition of claim 11 wherein the platinum is present in the amount of from 0.1 to 5 g/ft³ (3.5 to 177 g/m³) of the composition.
- 35 13. The catalyst composition of claim 12 wherein there is from 0.1 to 1.0 g/ft³ (3.5 to 35 g/m³) of platinum.
14. The catalyst composition of claim 8 further containing up to 0.5 g/ft³ (18 g/m³) of platinum dispersed on the catalytic material.
- 40 15. The catalyst composition of claim 10 or 14 wherein there is from 0.1 to 0.5 g/ft³ (3.5 to 18 g/m³) of platinum.
16. The catalyst composition of any of claims 8 to 15 wherein at least a catalytically effective amount of the platinum is dispersed on the ceria.
- 45 17. The catalyst composition of claim 16 wherein at least a catalytically effective amount of the platinum is dispersed on the alumina.
18. The catalyst composition of claim 8 including palladium dispersed on the catalytic material, the palladium being present in the amount of from 0.1 to 200 g/ft³ (3.5 to 7063 g/m³) of the composition.
- 50 19. The catalyst composition of claim 18 wherein the palladium is present in the amount of from 20 to 120 g/ft³ (706 to 4238 g/m³) of the composition.
- 55 20. The catalyst composition of any of claims 1 to 19 for use in the purification of a gas-borne stream, for example, a diesel engine exhaust, and for oxidising volatile organic components within the stream.
21. A method for oxidising oxidizable components of a gas-borne stream, which method comprises contacting the

stream with a catalyst composition at a temperature high enough to catalyze oxidation of at least some of the oxidizable components, the catalyst composition comprising a catalytic material consisting essentially of a combination of bulk ceria having a BET surface area of at least 10 m²/g and bulk alumina having a BET surface area of at least 10 m²/g.

- 5 22. A method according to claim 21 wherein the gas-borne stream comprises a diesel engine exhaust stream containing a volatile organic fraction, and wherein the stream is contacted with the catalyst composition at a temperature high enough to catalyze oxidation of at least some of the volatile organic fraction.
- 10 23. The method of claim 21 or claim 22 wherein each of the ceria and alumina comprises from 5 to 95 percent by weight of the combination.
24. The method of claim 21 or claim 22 wherein each of the ceria and alumina comprises from 10 to 90 percent by weight of the combination.
- 15 25. The method of claim 21 or claim 22 wherein each of the ceria and alumina comprises from 40 to 60 percent by weight of the combination.
26. The method of any of claims 21 to 25 wherein the temperature of the exhaust initially contacted with the catalyst composition is from 100°C to 800°C.
- 20 27. The method of any of claims 21 to 26 wherein the ceria and the alumina each has a BET surface area of from 25 m²/g to 200 m²/g.
- 25 28. The method of any of claims 21 to 27 wherein the ceria comprises aluminum-stabilized ceria.
29. The method of any of claims 21 to 28 wherein the ceria and the alumina are each disposed in respective discrete layers, one overlying the other.
- 30 30. The method of any of claims 21 to 29 wherein the catalyst composition further comprises catalytically active platinum or palladium dispersed on the catalytic material, the platinum or palladium being present in a moderate amount to limit or to reduce the oxidation of SO₂ to SO₃.
- 35 31. The method according to claim 30 wherein the catalytic material has dispersed thereon a catalytically effective amount of platinum not exceeding 25 g/ft³ (883 g/m³) of the catalyst composition.
32. The method of claim 30 further including platinum dispersed on the catalytic material, the platinum being present in the amount of from 0.1 to 15 g/ft³ (3.5 to 530 g/m³) of the composition.
- 40 33. The method of claim 32 wherein the platinum is present in the amount of from 0.1 to 5 g/ft³ (3.5 to 177 g/m³) of the composition.
34. The method of claim 33 wherein there is from 0.1 to 1.0 g/ft³ (3.5 to 35 g/m³) of platinum.
- 45 35. The method of claim 30 wherein the catalytic material has up to 0.5 g/ft³ (18 g/m³) of platinum dispersed thereon.
36. The method of claim 35 wherein the platinum is present in an amount of from 0.1 to 0.5 g/ft³ (3.5 to 18 g/m³) of the composition.
- 50 37. The method of claim 30 including palladium dispersed on the catalytic material, the palladium being present in the amount of from 0.1 to 200 g/ft³ (3.5 to 7063 g/m³) of the composition.
38. The method of claim 36 wherein the palladium is present in the amount of from 20 to 120 g/ft³ (706 to 4238 g/m³) of the composition.
- 55 39. The use, in a process for the catalytic oxidation of at least some of the volatile organic component of a diesel engine exhaust stream, of a catalyst to effect such oxidation whilst moderating the oxidation of SO₂ to SO₃, which catalyst comprises a refractory carrier on which is disposed a coating of a ceria-alumina catalytic material consisting essen-

tially of a combination of bulk ceria having a BET surface area of at least $10 \text{ m}^2/\text{g}$ and bulk alumina having a BET surface area of at least $10 \text{ m}^2/\text{g}$, the catalyst further comprising platinum dispersed on the said catalytic material, the platinum being present in an amount of from 0.1 to 5 g/ft^3 (3.5 to 177 g/m^3) of the said composition.

- 5 40. The catalyst composition of any of claims 1 to 8 comprising the refractory carrier, the ceria and the alumina in the absence of a platinum group metal.
41. The catalyst composition of any of claims 1 to 8 comprising the refractory material and the catalytic material consisting essentially of a combination of ceria and alumina, the catalyst further comprising up to 0.1 g/ft^3 (up to 3.5 g/m^3) of platinum dispersed on the catalytic material.

Patentansprüche

- 15 1. Eine Oxidationskatalysator-Zusammensetzung, umfassend einen hitzebeständigen Träger, auf dem eine Beschichtung aus einem katalytischen Ceroxid-Aluminiumoxid-Material aufgebracht ist, das im wesentlichen aus einer Kombination von Bulk-Ceroxid mit einer BET-Oberfläche von mindestens $10 \text{ m}^2/\text{g}$ und Bulk-Aluminiumoxid mit einer BET-Oberfläche von mindestens $10 \text{ m}^2/\text{g}$ besteht.
- 20 2. Die Katalysatorzusammensetzung nach Anspruch 1, worin das Ceroxid und Aluminiumoxid jeweils 5 bis 95 Gew.-% der Kombination umfassen.
3. Die Katalysatorzusammensetzung nach Anspruch 1, worin das Ceroxid und Aluminiumoxid jeweils 10 bis 90 Gew.-% der Kombination umfassen.
- 25 4. Die Katalysatorzusammensetzung nach Anspruch 1, worin das Ceroxid und Aluminiumoxid jeweils 40 bis 60 Gew.-% der Kombination umfassen.
5. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 4, worin das Ceroxid und das Aluminiumoxid jeweils in getrennten Schichten, eine über der anderen, angeordnet sind.
- 30 6. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 5, worin das Ceroxid ein Aluminium-stabilisiertes Ceroxid umfaßt.
7. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 6, worin das Ceroxid und das Aluminiumoxid jeweils eine BET-Oberfläche von $25 \text{ m}^2/\text{g}$ bis $200 \text{ m}^2/\text{g}$ haben.
- 35 8. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 7, die weiterhin katalytisch aktives Platin oder Palladium, dispergiert auf dem katalytischen Material, enthält, wobei das Platin oder Palladium in einer mäßigen Menge vorhanden ist, um die Oxidation von SO_2 zu SO_3 zu beschränken oder zu verringern.
- 40 9. Die Katalysatorzusammensetzung nach Anspruch 8, worin das katalytische Material darauf dispergiert eine katalytisch wirksame Menge von Platin hat, die 25 g/ft^3 (883 g/m^3) der Katalysatorzusammensetzung nicht übersteigt.
- 45 10. Die Katalysatorzusammensetzung nach Anspruch 8, worin das katalytische Material darauf dispergiert eine katalytisch wirksame Menge von Platin hat, die 15 g/ft^3 (530 g/m^3) der Katalysatorzusammensetzung nicht übersteigt.
11. Die Katalysatorzusammensetzung nach Anspruch 8, die weiterhin auf dem katalytischen Material dispergiertes Platin enthält, wobei das Platin in einer Menge von $0,1$ bis 15 g/ft^3 ($3,5$ bis 530 g/m^3) der Zusammensetzung vorhanden ist.
- 50 12. Die Katalysatorzusammensetzung nach Anspruch 11, worin das Platin in einer Menge von $0,1$ bis 5 g/ft^3 ($3,5$ bis 177 g/m^3) der Zusammensetzung vorhanden ist.
13. Die Katalysatorzusammensetzung nach Anspruch 12, worin $0,1$ bis $1,0 \text{ g/ft}^3$ ($3,5$ bis 35 g/m^3) Platin vorhanden ist.
- 55 14. Die Katalysatorzusammensetzung nach Anspruch 8, die weiterhin bis zu $0,5 \text{ g/ft}^3$ (18 g/m^3) Platin, dispergiert auf dem katalytischen Material, enthält.

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15. Die Katalysatorzusammensetzung nach Anspruch 10 oder 14, worin 0,1 bis 0,5 g/ft³ (3,5 bis 18 g/m³) Platin vorhanden ist.
- 5 16. Die Katalysatorzusammensetzung nach einem der Ansprüche 8 bis 15, worin mindestens eine katalytisch wirksame Menge des Platins auf dem Ceroxid dispergiert ist.
17. Die Katalysatorzusammensetzung nach Anspruch 16, worin mindestens eine katalytisch wirksame Menge des Platins auf dem Aluminiumoxid dispergiert ist.
- 10 18. Die Katalysatorzusammensetzung nach Anspruch 8, enthaltend Palladium, dispergiert auf dem katalytischen Material, wobei das Palladium in einer Menge von 0,1 bis 200 g/ft³ (3,5 bis 7063 g/m³) der Zusammensetzung vorhanden ist.
- 15 19. Die Katalysatorzusammensetzung nach Anspruch 18, worin das Palladium in einer Menge von 20 bis 120 g/ft³ (706 bis 4238 g/m³) der Zusammensetzung vorhanden ist.
- 20 20. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 19 zur Verwendung bei der Reinigung eines gashaltigen Stroms, z.B. eines Dieselmotorenabgases, und zur Oxidation flüchtiger, organischer Komponenten in dem Strom.
- 25 21. Ein Verfahren zum Oxidieren oxidierbarer Komponenten eines gashaltigen Stroms, umfassend den Kontakt des Stromes mit einer Katalysatorzusammensetzung bei einer Temperatur, die hoch genug ist, um die Oxidation von mindestens einem Teil der oxidierbaren Komponenten zu katalysieren, wobei die Katalysatorzusammensetzung ein katalytisches Material umfaßt, das im wesentlichen aus einer Kombination von Bulk-Ceroxid mit einer BET-Oberfläche von mindestens 10 m²/g und Bulk-Aluminiumoxid mit einer BET-Oberfläche von mindestens 10 m²/g besteht.
- 30 22. Ein Verfahren gemäß Anspruch 21, worin der gashaltige Strom einen Dieselmotoren-Abgasstrom umfaßt, der eine flüchtige, organische Fraktion enthält, und worin der Strom mit der Katalysatorzusammensetzung bei einer Temperatur in Kontakt gebracht wird, die hoch genug ist, um die Oxidation mindestens eines Teils der flüchtigen, organischen Fraktion zu katalysieren.
- 35 23. Das Verfahren nach Anspruch 21 oder 22, worin das Ceroxid und Aluminiumoxid jeweils 5 bis 95 Gew.-% der Kombination umfassen.
24. Verfahren nach Anspruch 21 oder 22, worin das Ceroxid und Aluminiumoxid jeweils 10 bis 90 Gew.-% der Kombination umfassen.
- 40 25. Das Verfahren nach Anspruch 21 oder 22, worin das Ceroxid und Aluminiumoxid jeweils 40 bis 60 Gew.-% der Kombination umfassen.
26. Das Verfahren nach einem der Ansprüche 1 bis 25, worin die Temperatur des am Anfang mit der Katalysatorzusammensetzung in Kontakt gebrachten Abgases 100 bis 800 °C beträgt.
- 45 27. Verfahren nach einem der Ansprüche 21 bis 26, worin das Ceroxid und das Aluminiumoxid jeweils eine BET-Oberfläche von 25 m²/g bis 200 m²/g haben.
28. Das Verfahren nach einem der Ansprüche 21 bis 27, worin das Ceroxid Aluminiumstabilisiertes Ceroxid umfaßt.
- 50 29. Das Verfahren nach einem der Ansprüche 21 bis 28, worin das Ceroxid und das Aluminiumoxid jeweils in getrennten Schichten, eine über der anderen, angeordnet sind.
30. Verfahren nach einem der Ansprüche 21 bis 29, worin die Katalysatorzusammensetzung weiterhin katalytisch aktives Platin oder Palladium, dispergiert auf dem katalytischen Material, umfaßt, wobei das Platin oder Palladium in einer mäßigen Menge vorhanden ist, um die Oxidation von SO₂ zu SO₃ zu beschränken oder zu verringern.
- 55 31. Das Verfahren nach Anspruch 30, worin das katalytische Material eine darauf dispergierte, katalytisch wirksame Menge von Platin enthält, die 25 g/ft³ (883 g/m³) der Katalysatorzusammensetzung nicht übersteigt.

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32. Das Verfahren nach Anspruch 30, weiter enthaltend Platin, dispergiert auf dem katalytischen Material, wobei das Platin in einer Menge von 0,1 bis 15 g/ft³ (3,5 bis 530 g/m³) der Zusammensetzung vorhanden ist.
- 5 33. Das Verfahren nach Anspruch 32, worin das Platin in einer Menge von 0,1 bis 5 g/ft³ (3,5 bis 177 g/m³) der Zusammensetzung vorhanden ist.
34. Das Verfahren nach Anspruch 33, worin 0,1 bis 1,0 g/ft³ (3,5 bis 35 g/m³) Platin vorhanden ist.
- 10 35. Das Verfahren nach Anspruch 30, worin das katalytische Material, darauf dispergiert, bis zu 0,5 g/ft³ (18 g/m³) Platin hat.
36. Das Verfahren nach Anspruch 35, worin das Platin in einer Menge von 0,1 bis 0,5 g/ft³ (3,5 bis 18 g/m³) der Zusammensetzung vorhanden ist.
- 15 37. Das Verfahren nach Anspruch 30, enthaltend Palladium, dispergiert auf dem katalytischen Material, wobei das Palladium in einer Menge von 0,1 bis 200 g/ft³ (3,5 bis 7063 g/m³) der Zusammensetzung vorhanden ist.
38. Das Verfahren nach Anspruch 37, worin das Palladium in einer Menge von 20 bis 120 g/ft³ (706 bis 4238 g/m³) der Zusammensetzung vorhanden ist.
- 20 39. Die Verwendung in einem Verfahren zur katalytischen Oxidation mindestens eines Teils der flüchtigen, organischen Komponente eines Dieselmotoren-Abgasstroms, eines Katalysators zum Bewirken einer solchen Oxidation, wobei die Oxidation von SO₂ zu SO₃ moderat erfolgt, wobei der Katalysator einen hitzebeständigen Träger umfaßt, auf dem eine Beschichtung eines katalytischen Ceroxid-Aluminiumoxid-Materials angeordnet ist, das im wesentlichen aus einer Kombination von Bulk-Ceroxid mit einer BET-Oberfläche von mindestens 10 m²/g und Bulk-Aluminiumoxid mit einer BET-Oberfläche von mindestens 10 m²/g besteht, und der Katalysator weiterhin auf dem katalytischen Material dispergiertes Platin enthält, wobei das Platin in einer Menge von 0,1 bis 5 g/ft³ (3,5 bis 177 g/m³) der Zusammensetzung vorhanden ist.
- 25 40. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 8, umfassend den hitzebeständigen Träger, das Ceroxid und das Aluminiumoxid, in Abwesenheit eines Metalls der Platingruppe.
- 30 41. Die Katalysatorzusammensetzung nach einem der Ansprüche 1 bis 8, umfassend das hitzebeständige Material und das katalytische Material, bestehend im wesentlichen aus einer Kombination von Ceroxid und Aluminiumoxid, wobei der Katalysator weiterhin bis zu 0,1 g/ft³ (bis zu 3,5 g/m³) Platin, dispergiert auf dem katalytischen Material, umfaßt.
- 35

Revendications

- 40 1. Composition de catalyseur d'oxydation qui comprend un support réfractaire sur lequel est disposé un revêtement d'un matériau catalytique d'oxyde cérique-alumine consistant essentiellement en une combinaison d'oxyde cérique massif ayant une surface spécifique BET d'au moins 10 m²/g et d'alumine massive ayant une surface spécifique BET de 10 m²/g.
- 45 2. Composition de catalyseur de la revendication 1, dans laquelle l'oxyde cérique et l'alumine constituent chacun 5 à 95 pour cent en poids de la combinaison.
3. Composition de catalyseur de la revendication 1, dans laquelle l'oxyde cérique et l'alumine constituent chacun 10 à 90 pour cent en poids de la combinaison.
- 50 4. Composition de catalyseur de la revendication 1, dans laquelle l'oxyde cérique et l'alumine constituent chacun 40 à 60 pour cent en poids de la combinaison.
- 55 5. Composition de catalyseur de l'une quelconque des revendications 1 à 4, dans laquelle l'oxyde cérique et l'alumine sont chacun disposés en couches discrètes respectives, l'une superposée à l'autre.
6. Composition de catalyseur de l'une quelconque des revendications 1 à 5, dans laquelle l'oxyde cérique comprend un oxyde cérique stabilisé par de l'aluminium.

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7. Composition de catalyseur de l'une quelconque des revendications 1 à 6, dans laquelle l'oxyde cérique et l'alumine ont chacun une surface spécifique BET comprise entre 25 m²/g et 200 m²/g.
- 5 8. Composition de catalyseur de l'une quelconque des revendications 1 à 7, comprenant en outre du platine ou du palladium catalytiquement actif dispersé sur le matériau catalytique, le platine ou le palladium étant présents en une quantité modérée permettant de limiter ou de réduire l'oxydation de SO₂ en SO₃.
9. Composition de catalyseur de la revendication 8, dans laquelle le matériau catalytique porte dispersé sur lui une quantité efficace de platine ne dépassant pas 25 g/pied³ (883 g/m³) de la composition de catalyseur.
- 10 10. Composition de catalyseur de la revendication 8, dans laquelle le matériau catalytique porte dispersé sur lui une quantité efficace de platine ne dépassant pas 15 g/pied³ (530 g/m³) de la composition de catalyseur.
11. Composition de catalyseur de la revendication 8, comprenant en outre du platine dispersé sur le matériau catalytique, le platine étant présent dans la quantité allant de 0,1 à 15 g/pied³ (3,5 à 530 g/m³) de la composition.
- 15 12. Composition de catalyseur de la revendication 11, dans laquelle le platine est présent dans la quantité allant de 0,1 à 5 g/pied³ (3,5 à 177 g/m³) de la composition.
- 20 13. Composition de catalyseur de la revendication 12, dans laquelle il y a de 0,1 à 1,0 g/pied³ (3,5 à 35 g/pied³) de platine.
14. Composition de catalyseur de la revendication 8, comprenant en outre jusqu'à 0,5 g/pied³ (18 g/m³) de platine dispersé sur le matériau catalytique.
- 25 15. Composition de catalyseur de la revendication 10 ou 14, dans laquelle il y a de 0,1 à 0,5 g/pied³ (3,5 à 18 g/m³) de platine.
16. Composition de catalyseur de l'une quelconque des revendications 8 à 15, dans laquelle au moins une quantité catalytiquement efficace du platine est dispersée sur l'oxyde cérique.
- 30 17. Composition de catalyseur de la revendication 16, dans laquelle au moins une quantité catalytiquement efficace du platine est dispersée sur l'alumine.
- 35 18. Composition de catalyseur de la revendication 8 comprenant du palladium dispersé sur le matériau catalytique, le palladium étant présent dans la quantité allant de 0,1 à 200 g/pied³ (3,5 à 7063 g/m³) de la composition.
19. Composition de catalyseur de la revendication 18, dans laquelle le palladium est présent dans la quantité allant de 20 à 120 g/pied³ (706 à 4238 g/m³) de la composition.
- 40 20. Composition de catalyseur de l'une quelconque des revendications 1 à 19, destinée à être utilisée dans la purification d'un flux en suspension dans un gaz, par exemple, un échappement de moteur Diesel, et destinée à l'oxydation de composants organiques volatils à l'intérieur du flux.
- 45 21. Méthode d'oxydation de composants oxydables d'un flux en suspension dans un gaz, cette méthode comprenant les étapes consistant à mettre en contact le flux avec une composition de catalyseur à une température suffisamment élevée pour catalyser l'oxydation d'au moins certains des composants oxydables, la composition de catalyseur comprenant un matériau catalytique consistant essentiellement en une combinaison d'oxyde cérique massif ayant une surface spécifique BET d'au moins 10 m²/g et d'alumine massif ayant une surface spécifique BET d'au moins 10 m²/g.
- 50 22. Méthode selon la revendication 21, dans laquelle le flux en suspension dans un gaz comprend un flux d'échappement de moteur Diesel contenant une fraction organique volatile, et dans laquelle le flux est mis en contact avec la composition de catalyseur à une température suffisamment élevée pour catalyser l'oxydation d'au moins une partie de la fraction organique volatile.
- 55 23. Méthode de la revendication 21 ou de la revendication 22, dans laquelle l'oxyde cérique et l'alumine constituent chacun 5 à 95 pour cent en poids de la combinaison.

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24. Méthode de la revendication 21 ou de la revendication 22, dans laquelle l'oxyde cérique et l'alumine constituent chacun 10 à 90 pour cent en poids de la combinaison.
- 5 25. Méthode de la revendication 21 ou de la revendication 22, dans laquelle l'oxyde cérique et l'alumine constituent chacun 40 à 60 pour cent en poids de la combinaison.
26. Méthode de l'une quelconque des revendications 21 à 25, dans laquelle la température de l'échappement initialement mis en contact avec la composition de catalyseur est comprise entre 100°C et 800°C.
- 10 27. Méthode de l'une quelconque des revendications 21 à 26, dans laquelle l'oxyde cérique et l'alumine ont chacun une surface spécifique BET comprise entre 25 m²/g à 200 m²/g.
28. Méthode de l'une quelconque des revendications 21 à 27, dans laquelle l'oxyde cérique comprend de l'oxyde cérique stabilisé par de l'aluminium.
- 15 29. Méthode de l'une quelconque des revendications 21 à 28, dans laquelle l'oxyde cérique et l'alumine sont chacun disposés en couches discrètes respectives, l'une superposée à l'autre.
30. Méthode selon l'une quelconque des revendications 21 à 29, dans laquelle la composition de catalyseur comprend en outre du platine ou du palladium catalytiquement actif dispersé sur le matériau catalytique, le platine ou le palladium étant présents en une quantité modérée permettant de limiter ou de réduire l'oxydation de SO₂ en SO₃.
- 20 31. Méthode selon la revendication 30, dans laquelle le matériau catalytique porte dispersé sur lui une quantité efficace de platine ne dépassant pas 25 g/pied³ (883 g/m³) de la composition de catalyseur.
- 25 32. Méthode de la revendication 30, comprenant en outre du platine dispersé sur le matériau catalytique, le platine étant présent dans la quantité allant de 0,1 à 15 g/pied³ (3,5 à 530 g/m³) de la composition.
- 30 33. Méthode de la revendication 32, dans laquelle le platine est présent dans la quantité allant de 0,1 à 5 g/pied³ (3,5 à 177 g/m³) de la composition.
34. Méthode de la revendication 33, dans laquelle il y a de 0,1 à 1,0 g/pied³ (3,5 à 35 g/m³) de platine.
- 35 35. Méthode de la revendication 30, dans laquelle le matériau catalytique comprend jusqu'à 0,5 g/pied³ (18 g/m³) de platine dispersé dessus.
36. Méthode de la revendication 32, dans laquelle le platine est présent dans la quantité allant de 0,1 à 0,5 g/pied³ (3,5 à 18 g/m³) de la composition.
- 40 37. Méthode de la revendication 30, incluant du palladium dispersé sur le matériau catalytique, le palladium étant présent dans la quantité allant de 0,1 à 200 g/pied³ (3,5 à 7063 g/m³) de la composition.
- 45 38. Méthode de la revendication 36, dans laquelle le palladium est présent dans la quantité allant de 20 à 120 g/pied³ (706 à 4238 g/m³) de la composition.
- 50 39. Utilisation, dans un procédé destiné à l'oxydation d'au moins certains composants organiques volatils d'un flux d'échappement de moteur Diesel, d'un catalyseur permettant d'effectuer cette oxydation tout en modérant l'oxydation de SO₂ en SO₃, ce catalyseur comprenant un support réfractaire sur lequel est disposé un revêtement d'un matériau catalytique d'oxyde cérique-alumine consistant essentiellement en une combinaison d'oxyde cérique massif ayant une surface spécifique BET d'au moins 10 m²/g et d'alumine massive ayant une surface spécifique de 10 m²/g, le catalyseur comprenant en outre du platine dispersé sur le matériau catalytique, le platine étant présent dans la quantité allant de 0,1 à 5 g/pied³ (3,5 à 177 g/m³) de ladite composition.
- 55 40. Composition de catalyseur de l'une quelconque des revendications 1 à 8, comprenant le support réfractaire, l'oxyde cérique et l'alumine en l'absence d'un métal du groupe du platine.
41. Composition de catalyseur de l'une quelconque des revendications 1 à 8, comprenant le matériau réfractaire et le matériau catalytique consistant essentiellement en une combinaison d'oxyde cérique et d'alumine, le catalyseur

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comprenant en outre jusqu'à 0,1 g/ft³ (jusqu'à 3,5 g/m³) de platine dispersé sur le matériau catalytique.

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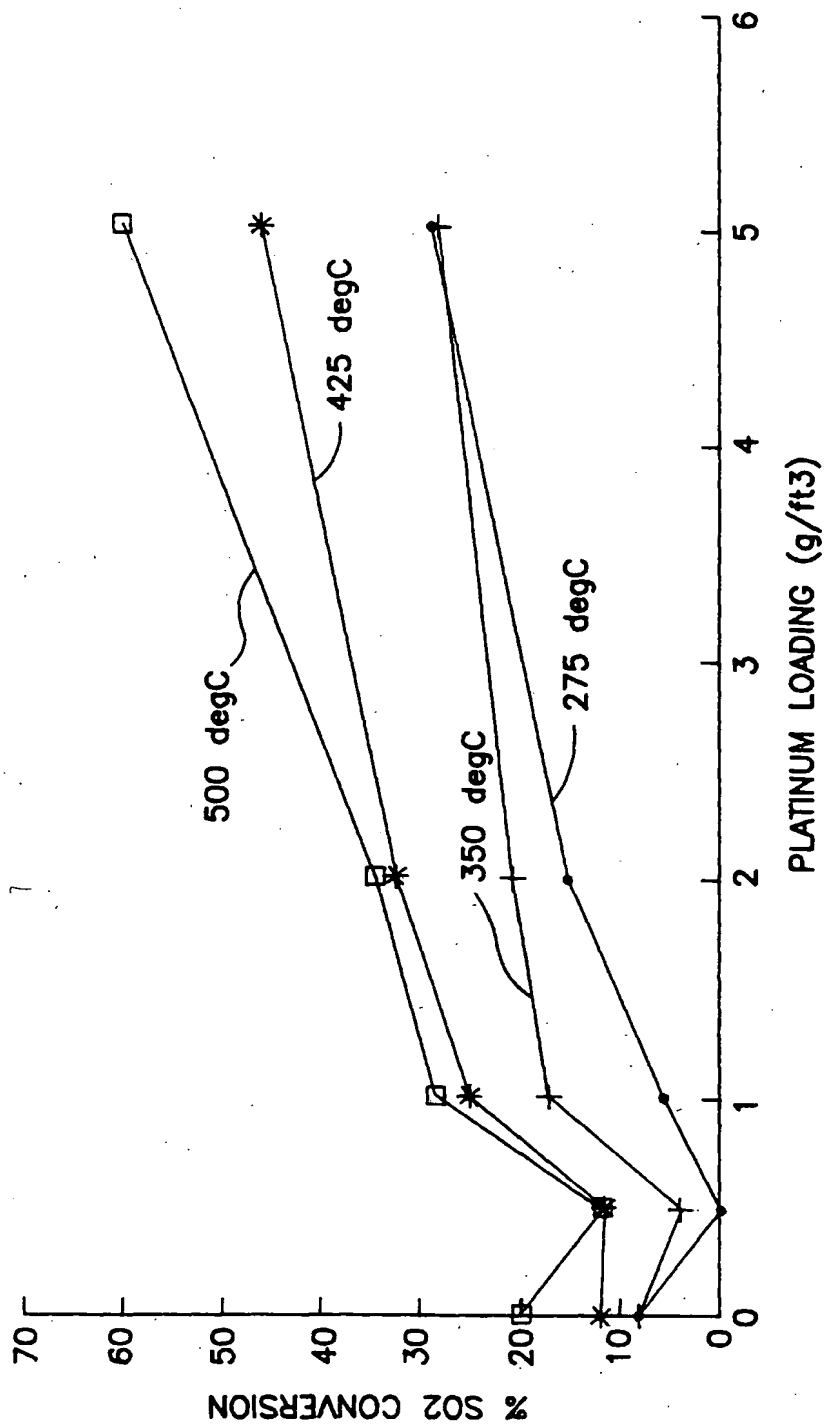


FIG. 1

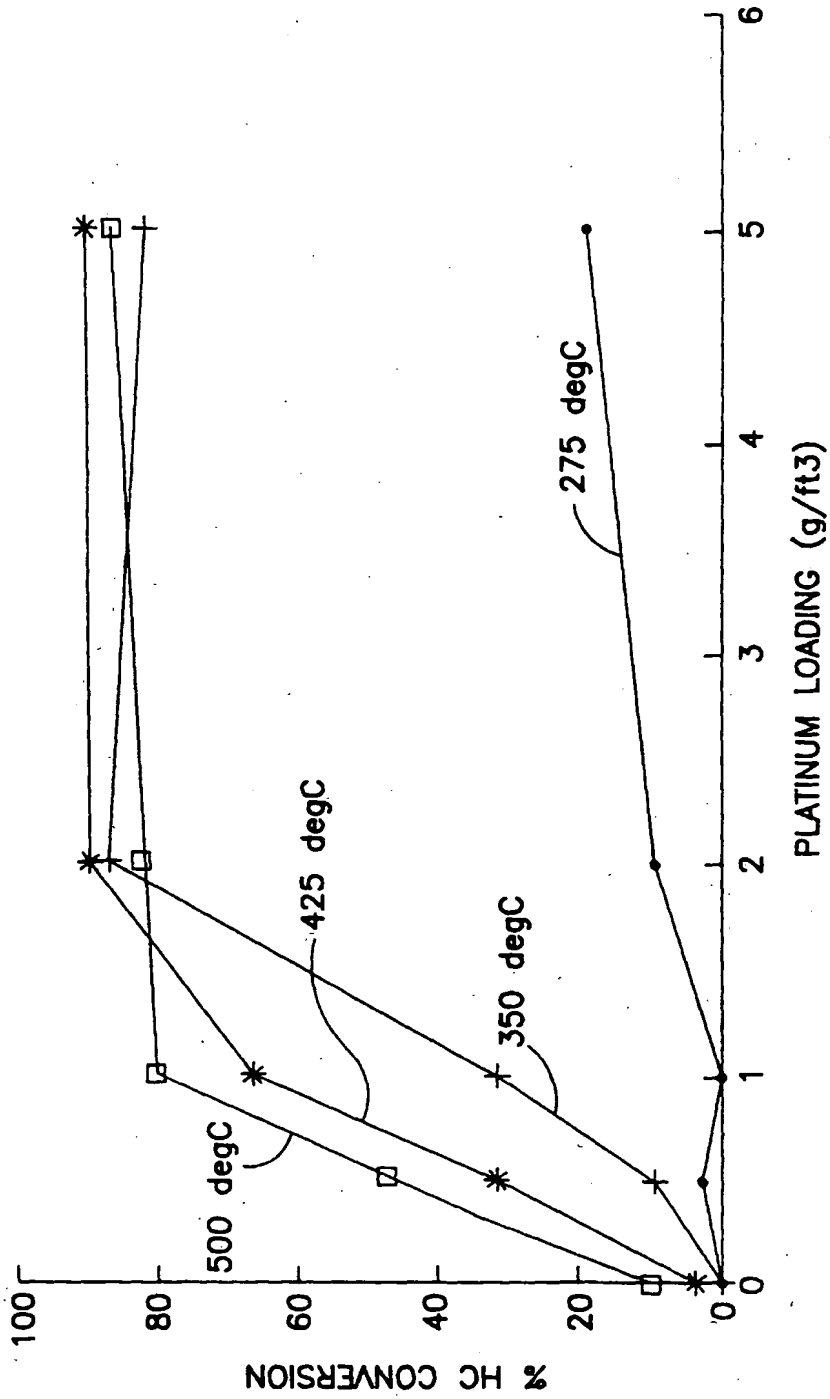
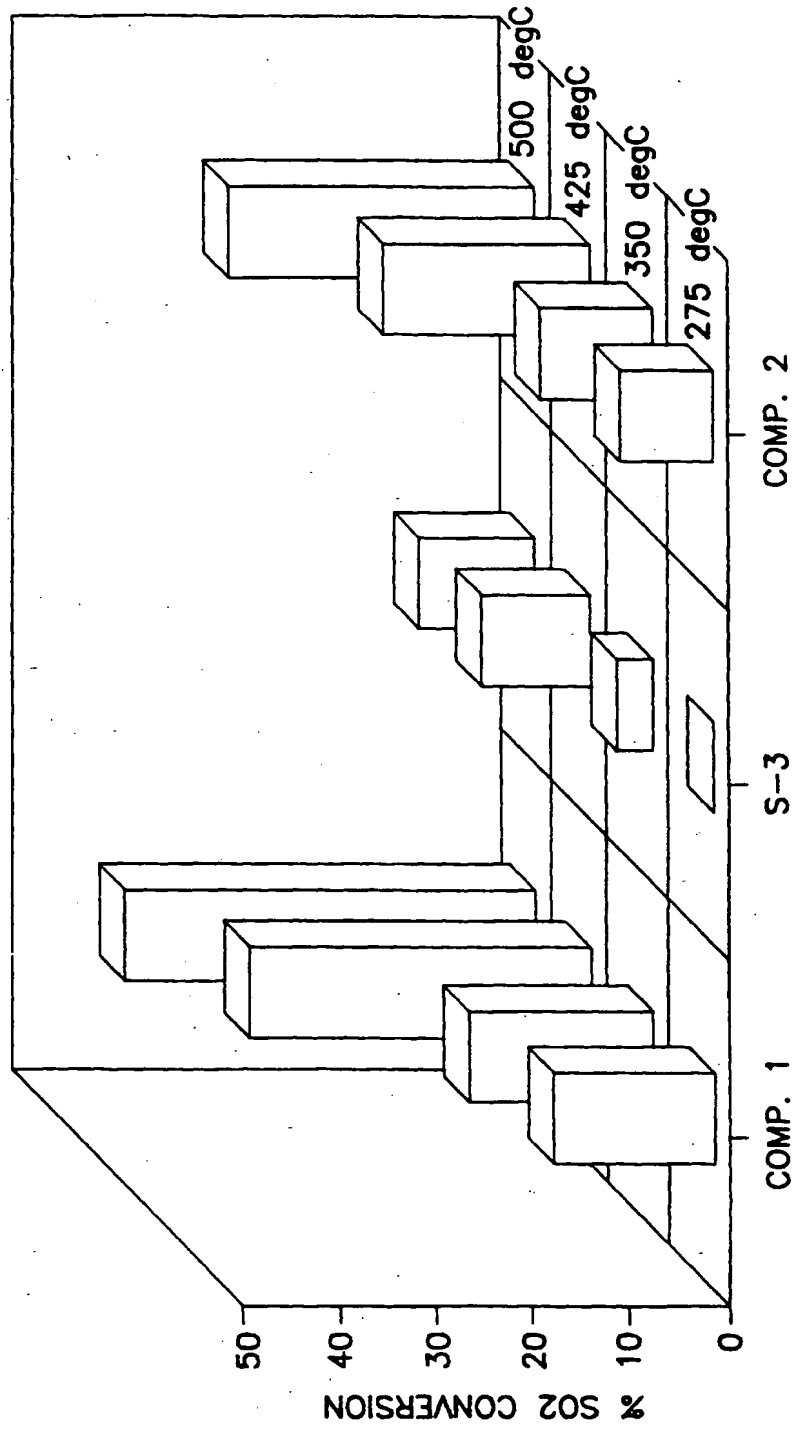


FIG. 2

FIG. 3



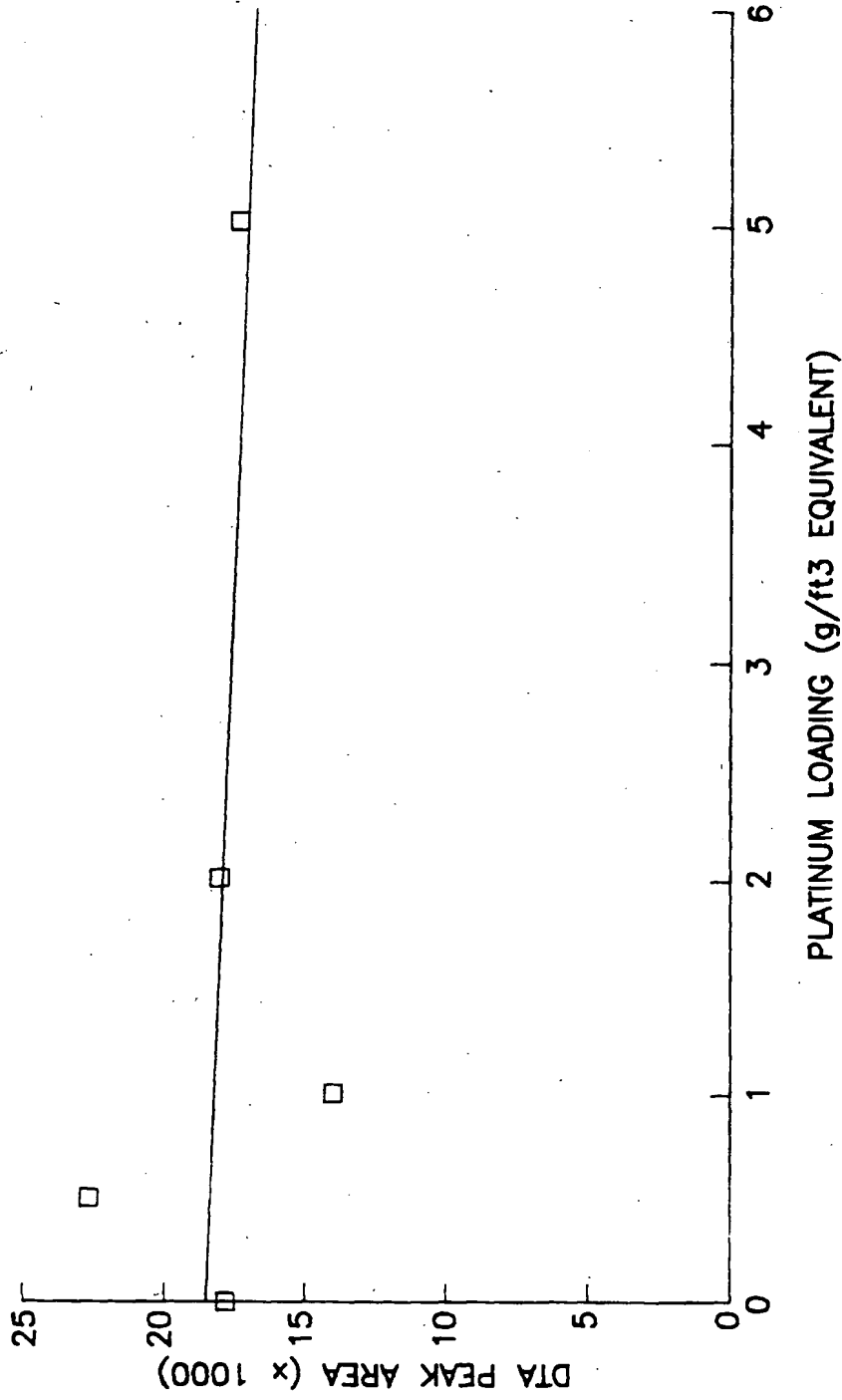


FIG. 4

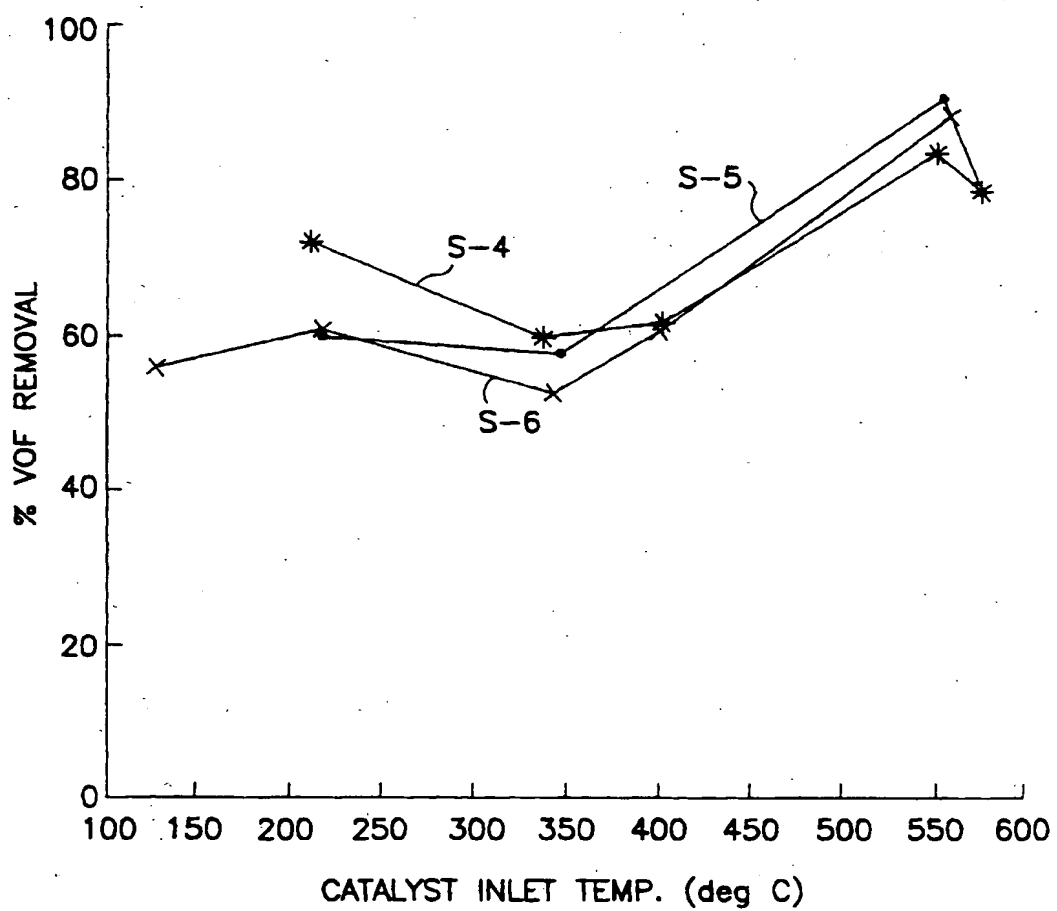


FIG. 5

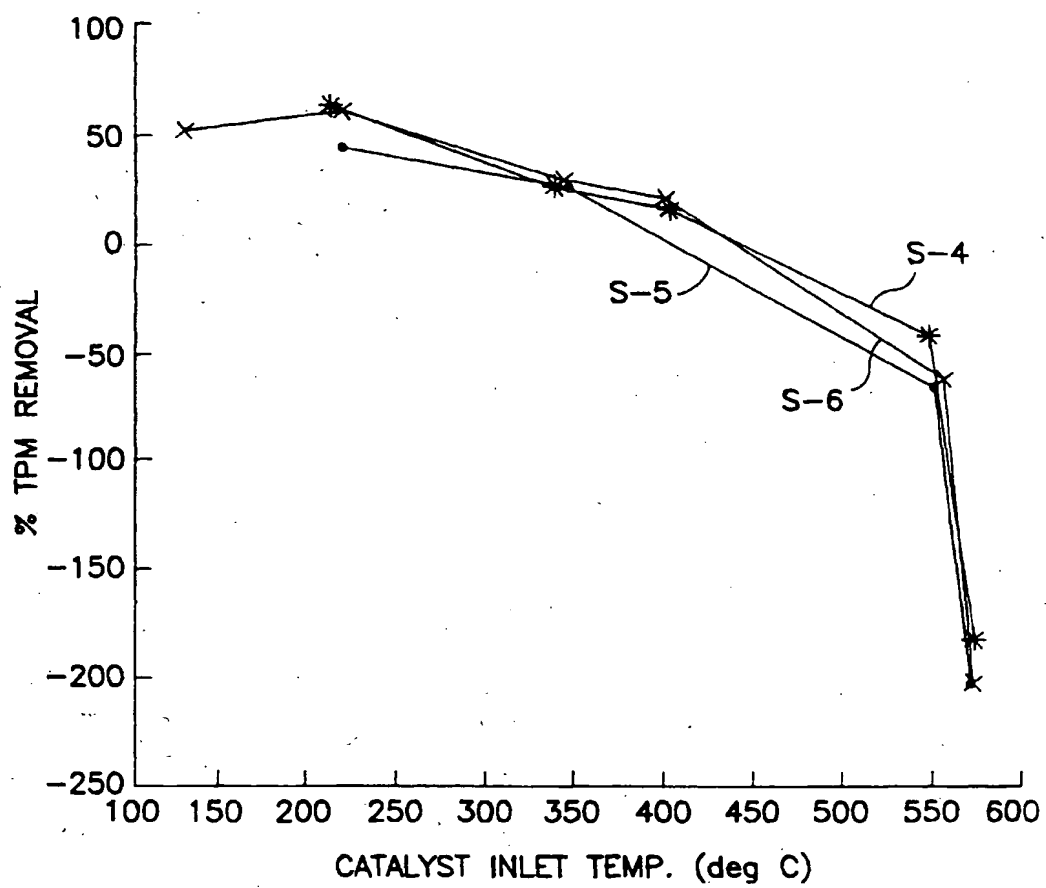


FIG. 6

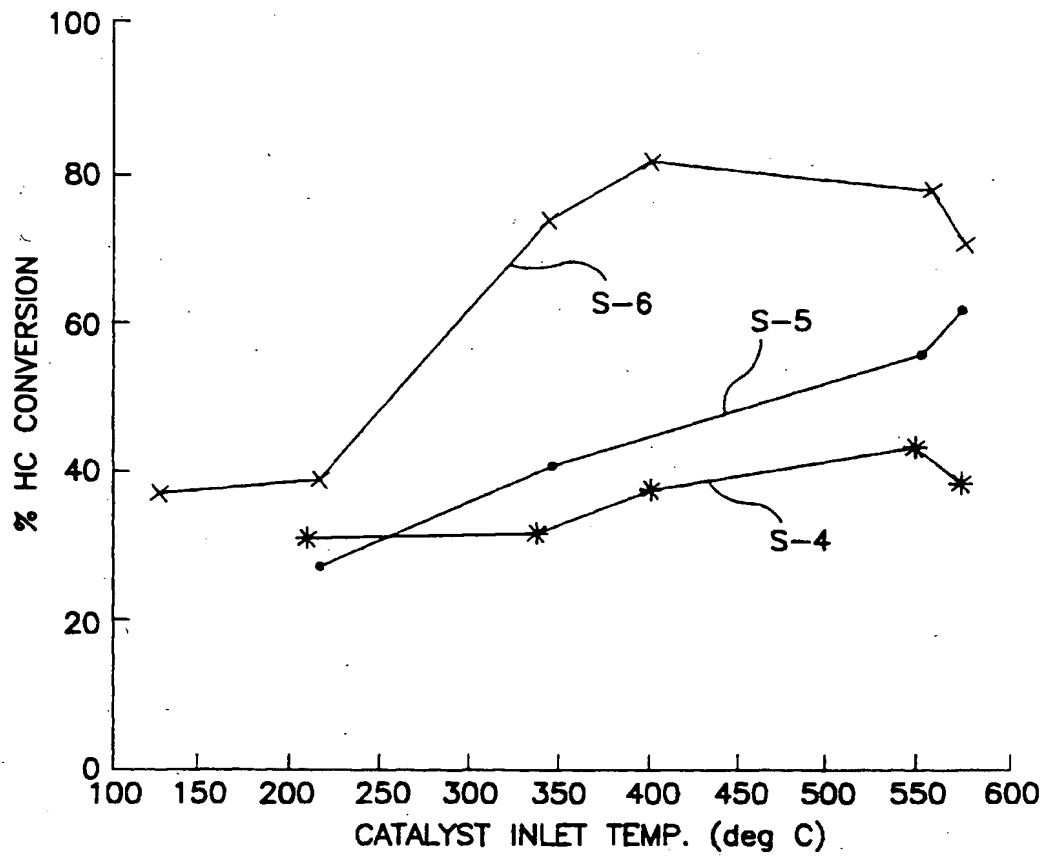


FIG. 7

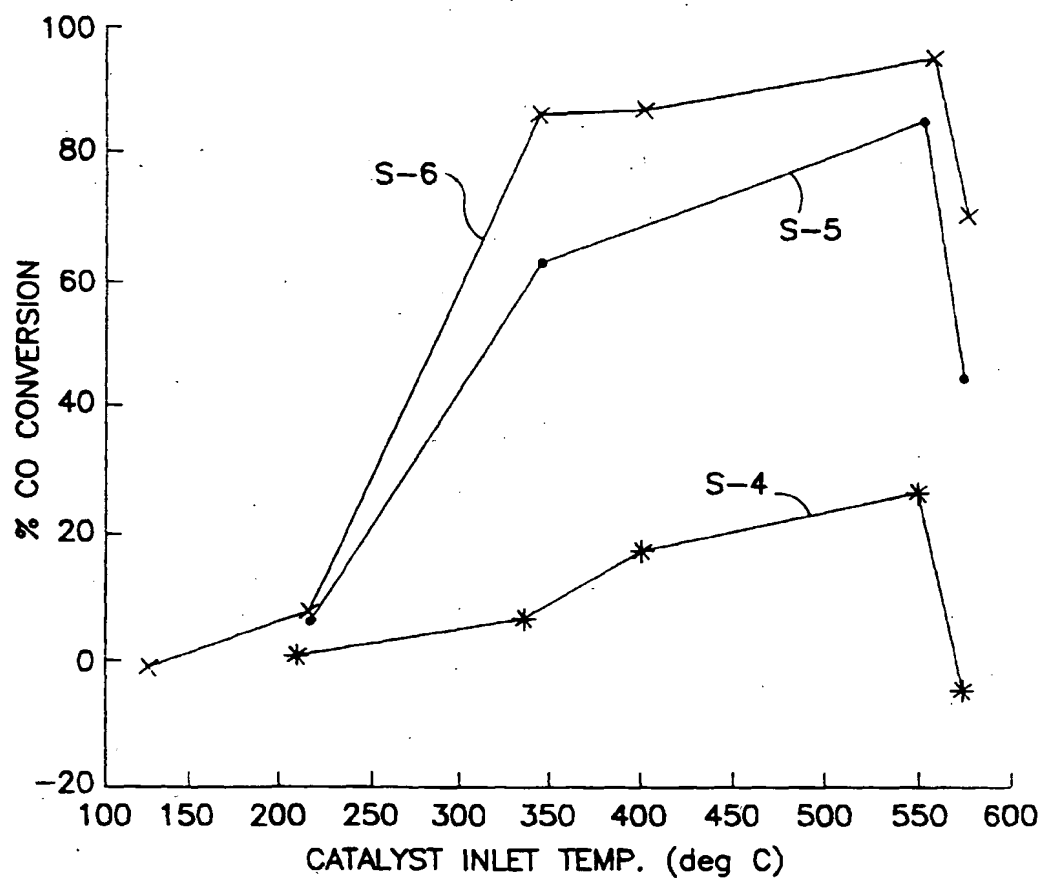


FIG. 8